



Ch. 4. Atomic Structures



Niels Henrik David Bohr
(1885-1962)



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Thomson



Joseph John Thomson (1856-1940)

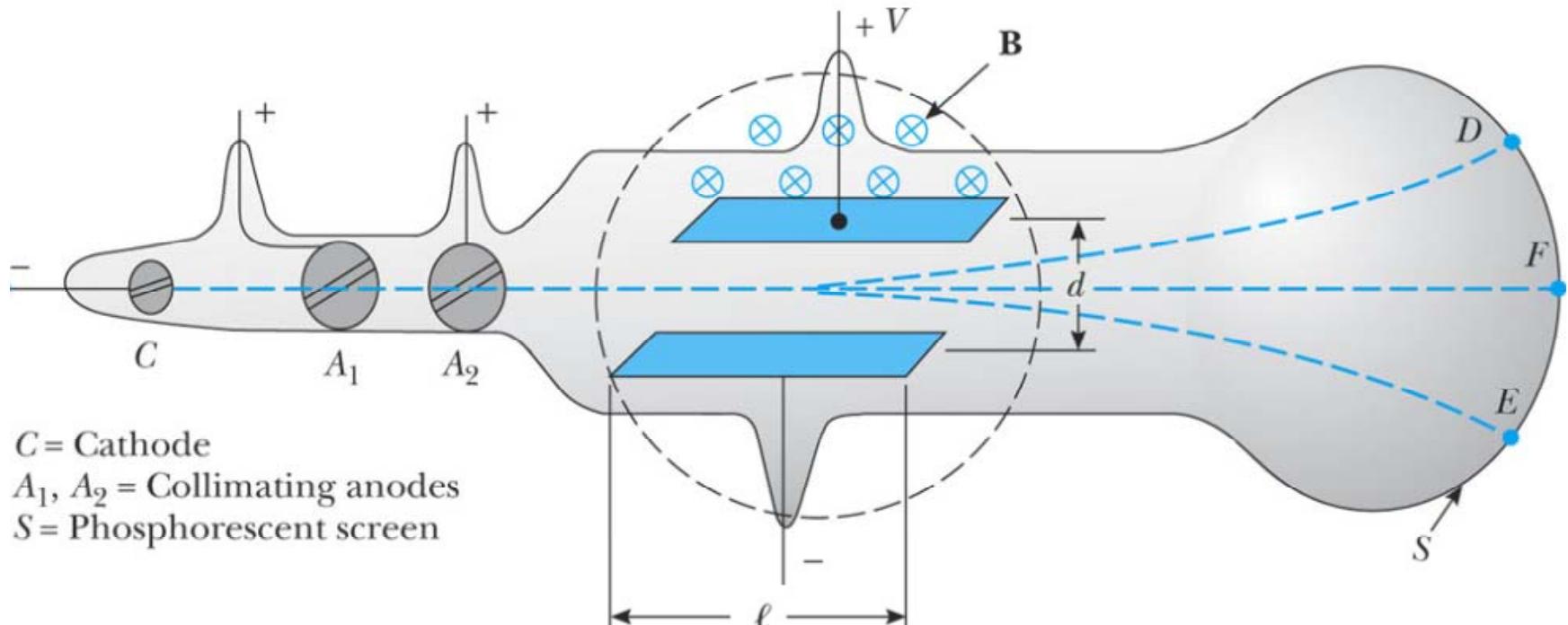
<http://nobelprize.org/physics/laureates/1906/thomson-bio.html>

... he achieved the most brilliant work of his life - an original study of cathode rays culminating in the discovery of the electron, which was announced during the course of his evening lecture to the Royal Institution on Friday, April 30, 1897.



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C = Cathode

A_1, A_2 = Collimating anodes

S = Phosphorescent screen

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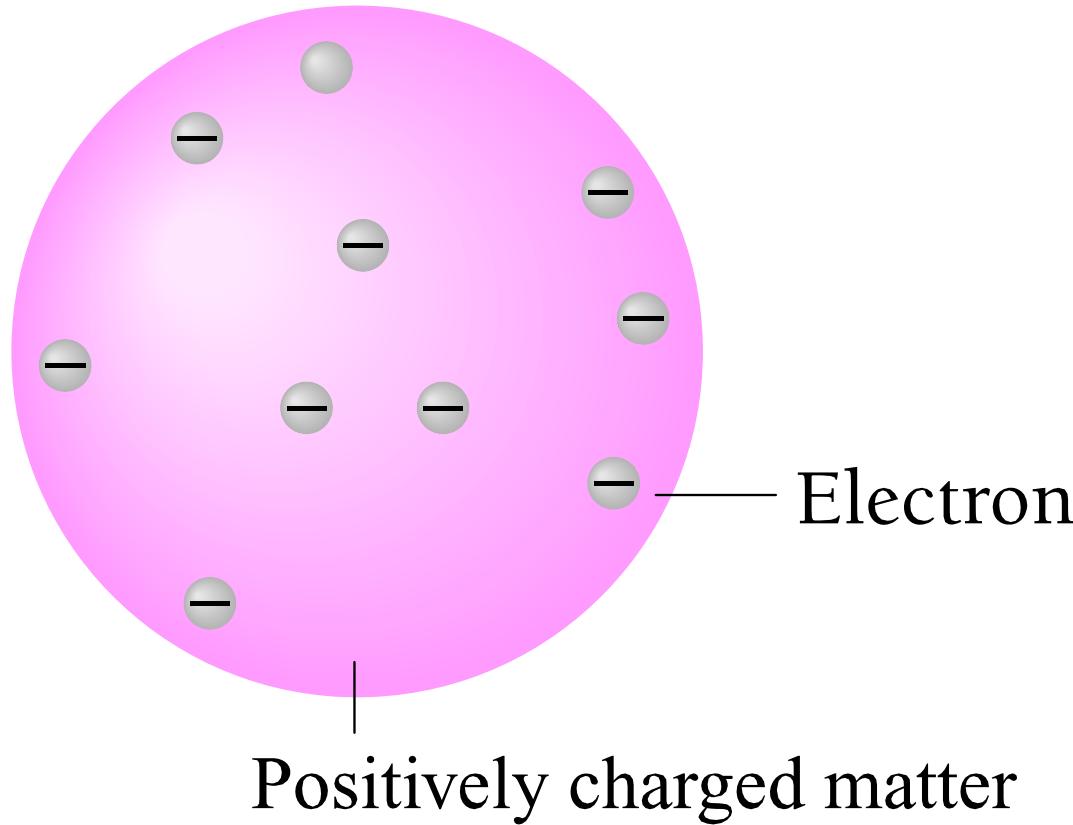


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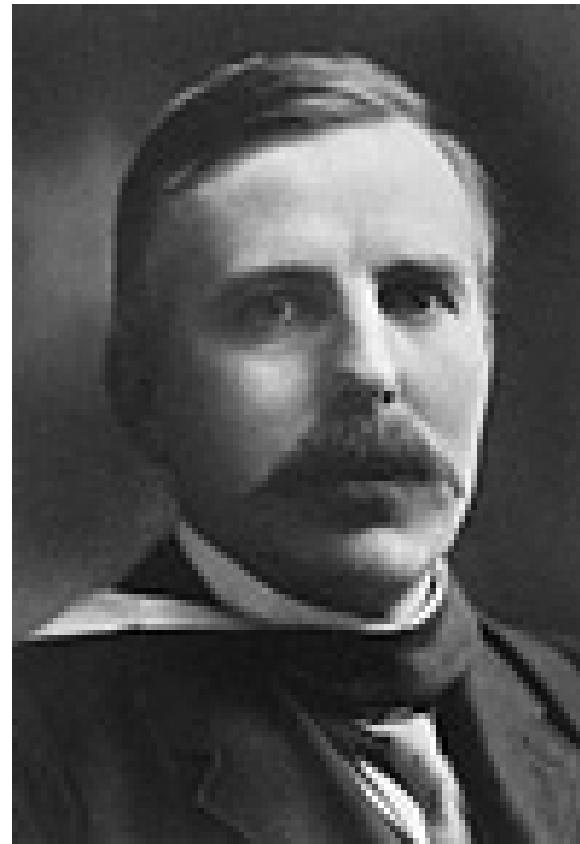


Thomson의 원자 모델





Rutherford



Ernest Rutherford (1871-1937)



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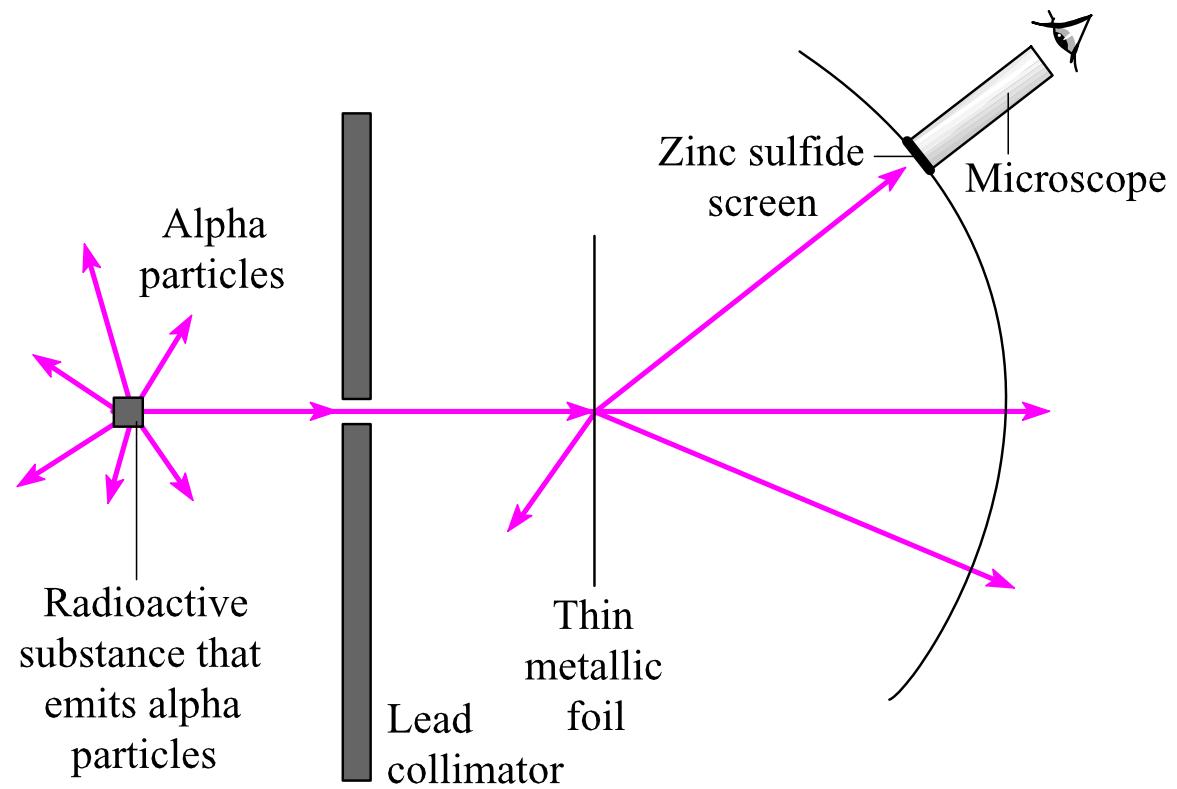
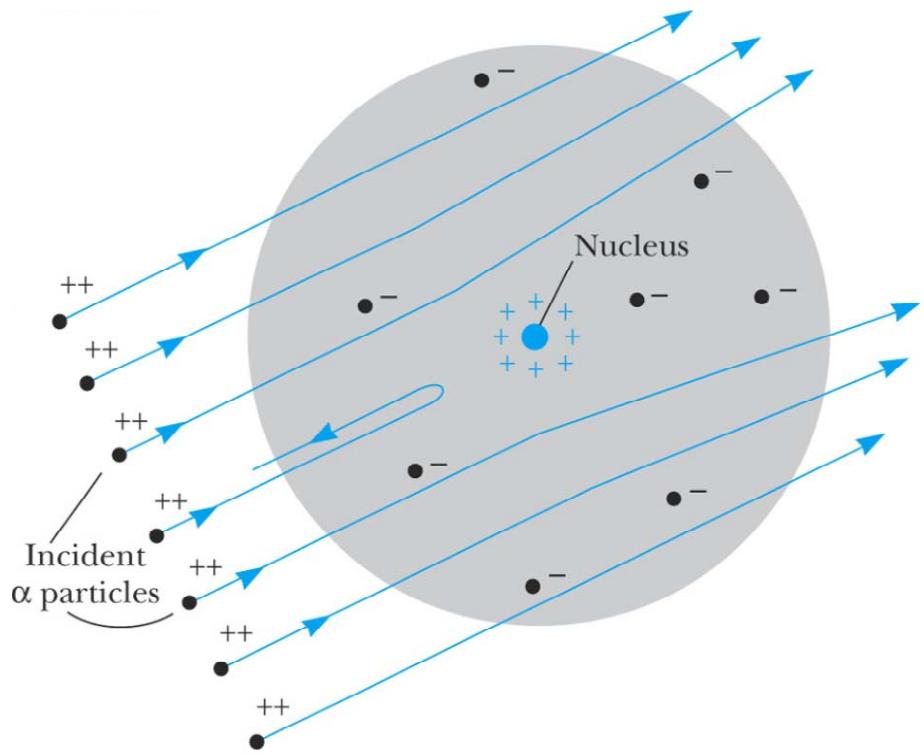
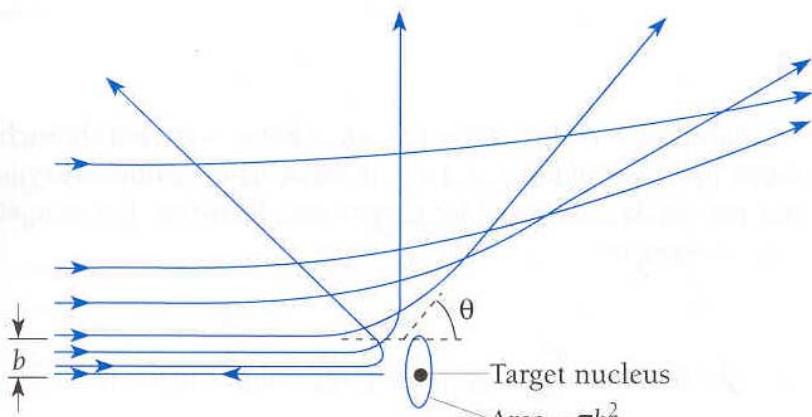


그림 4.2 Rutherford의 산란 실험



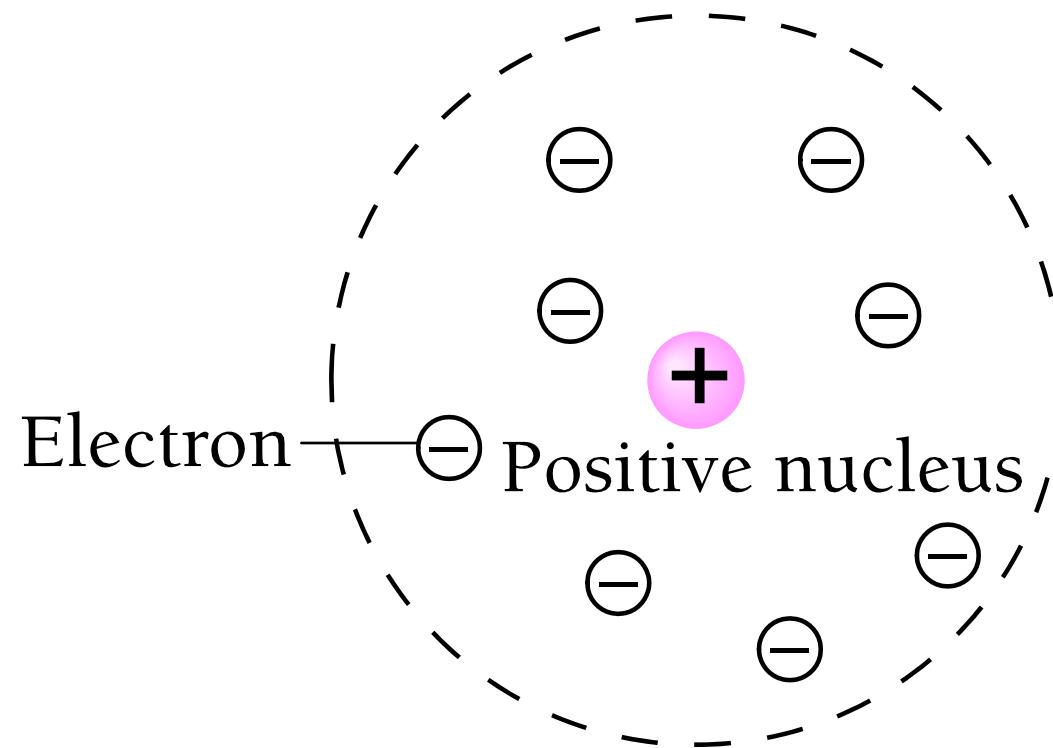


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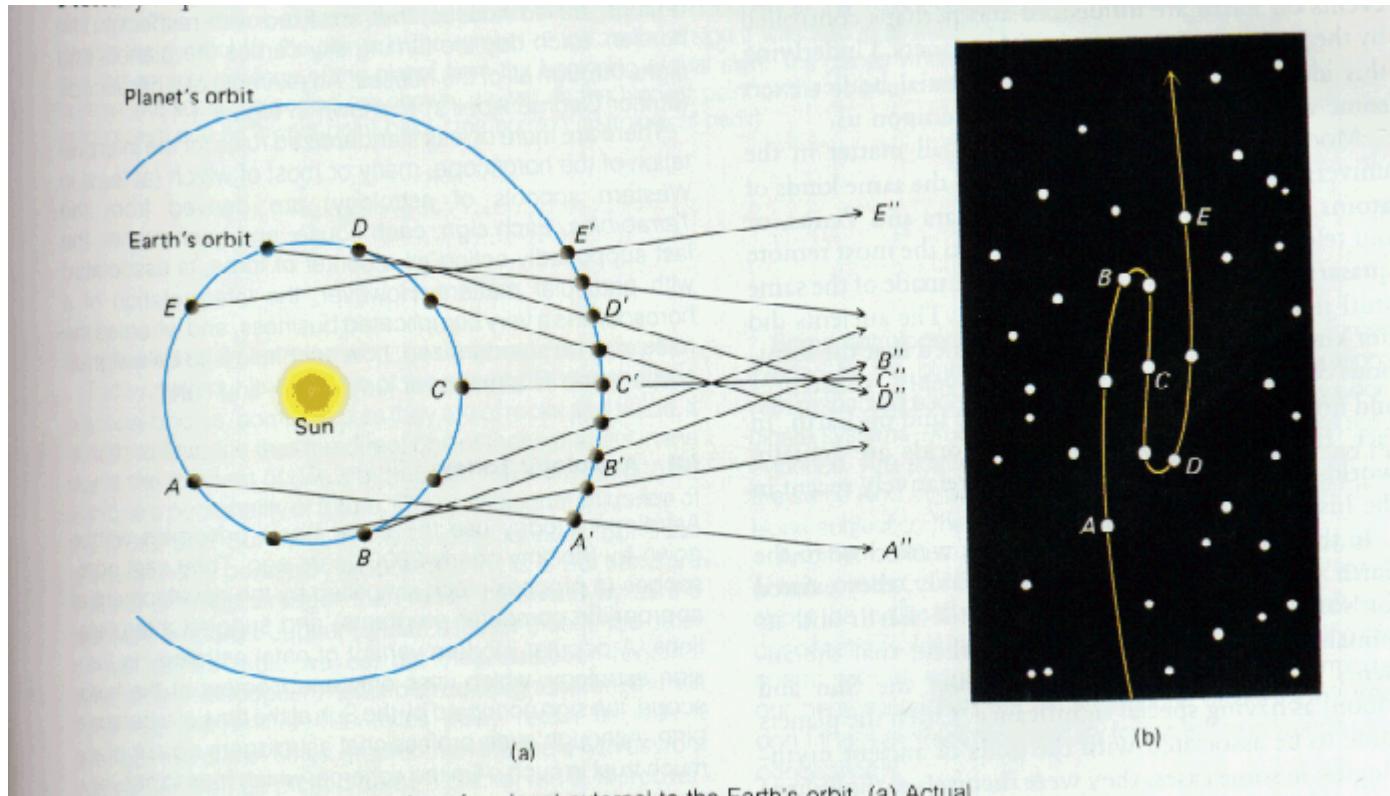


Rutherford의 원자 모델



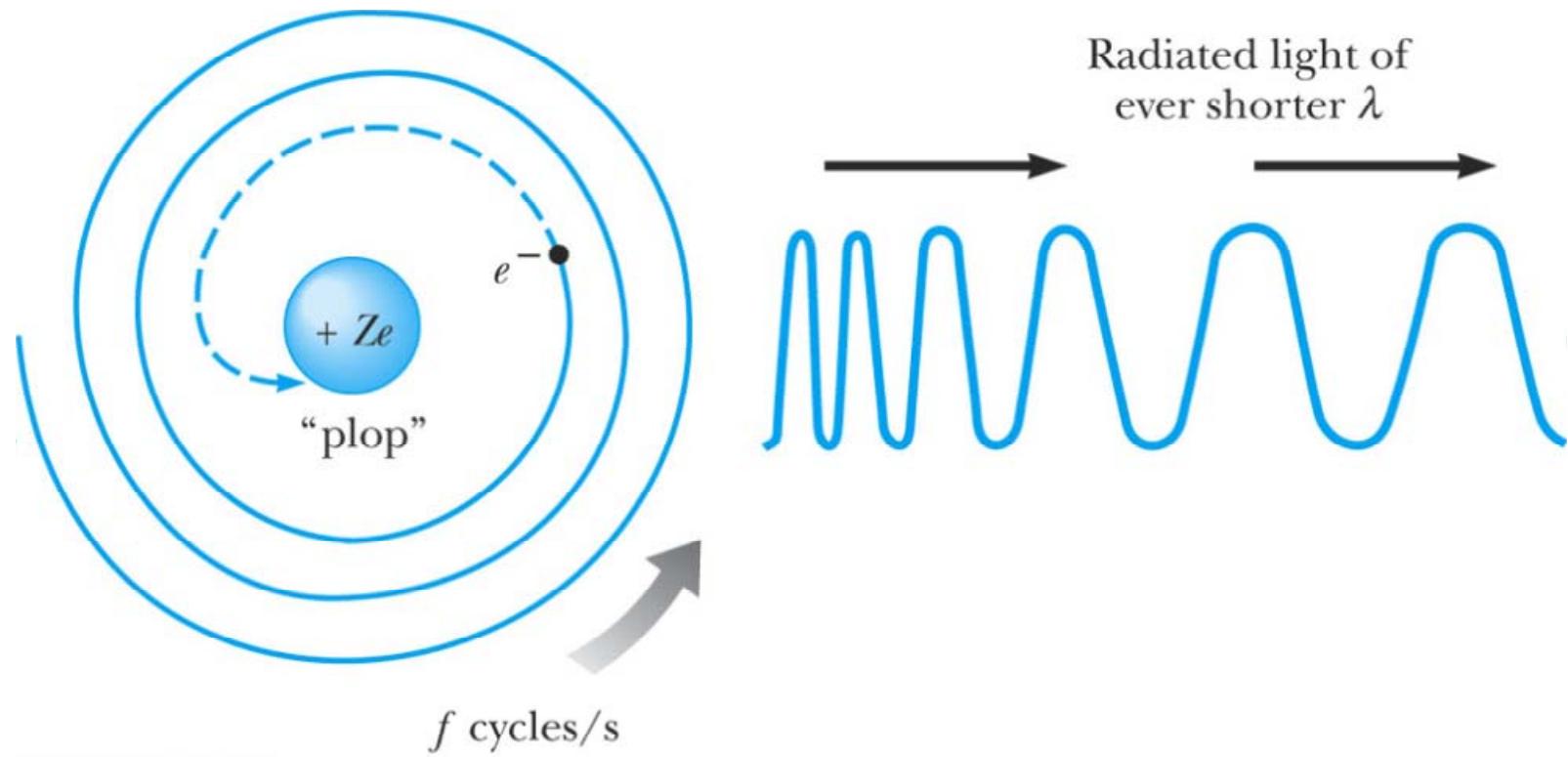


태양과 행성 – No Problem





Dilemma



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왜? - Maxwell's Equations



$$\vec{\nabla} \cdot \vec{D} = \rho$$

$$\vec{\nabla} \cdot \vec{B} = 0$$

$$\vec{\nabla} \times \vec{E} = -\frac{\partial \vec{B}}{\partial t}$$

$$\vec{\nabla} \times \vec{H} = \vec{J} + \frac{\partial \vec{D}}{\partial t}$$

$$\text{cf. } \vec{D} = \epsilon \vec{E}, \quad \vec{B} = \mu \vec{H}$$



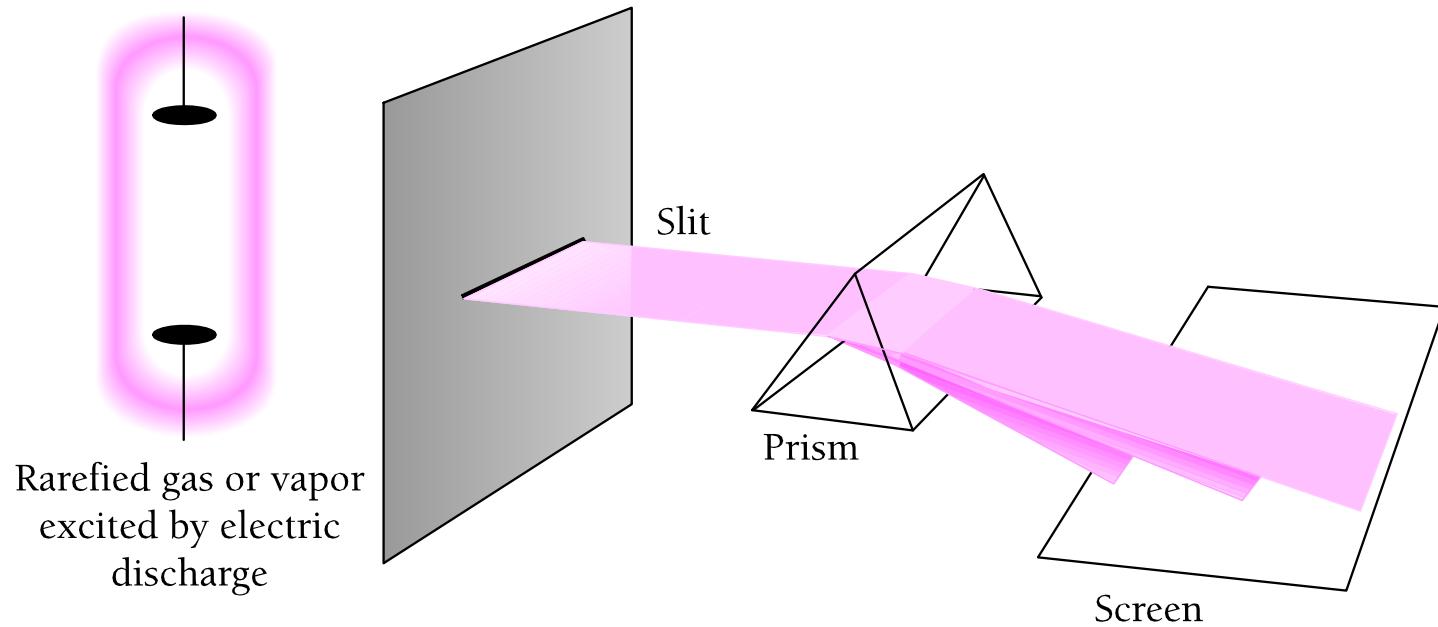
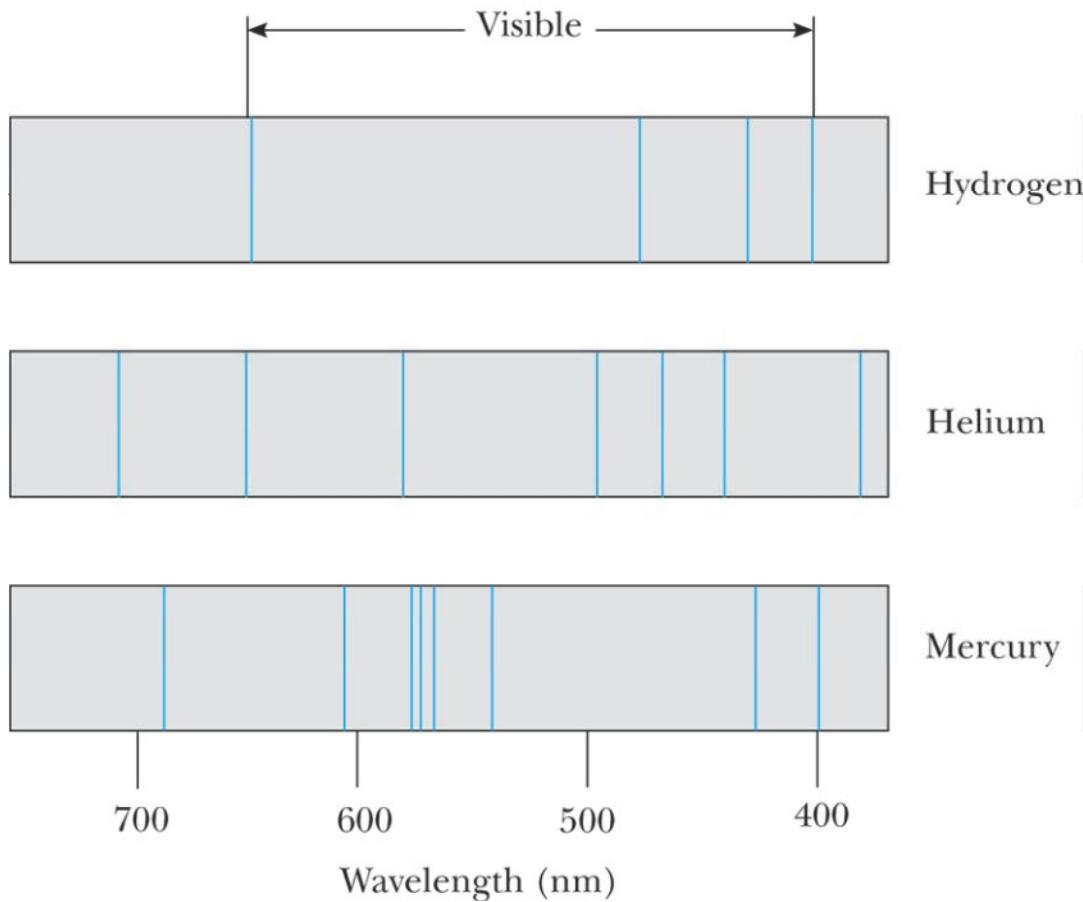


그림 4.7 이상적인 분광기.





Oh, My God...



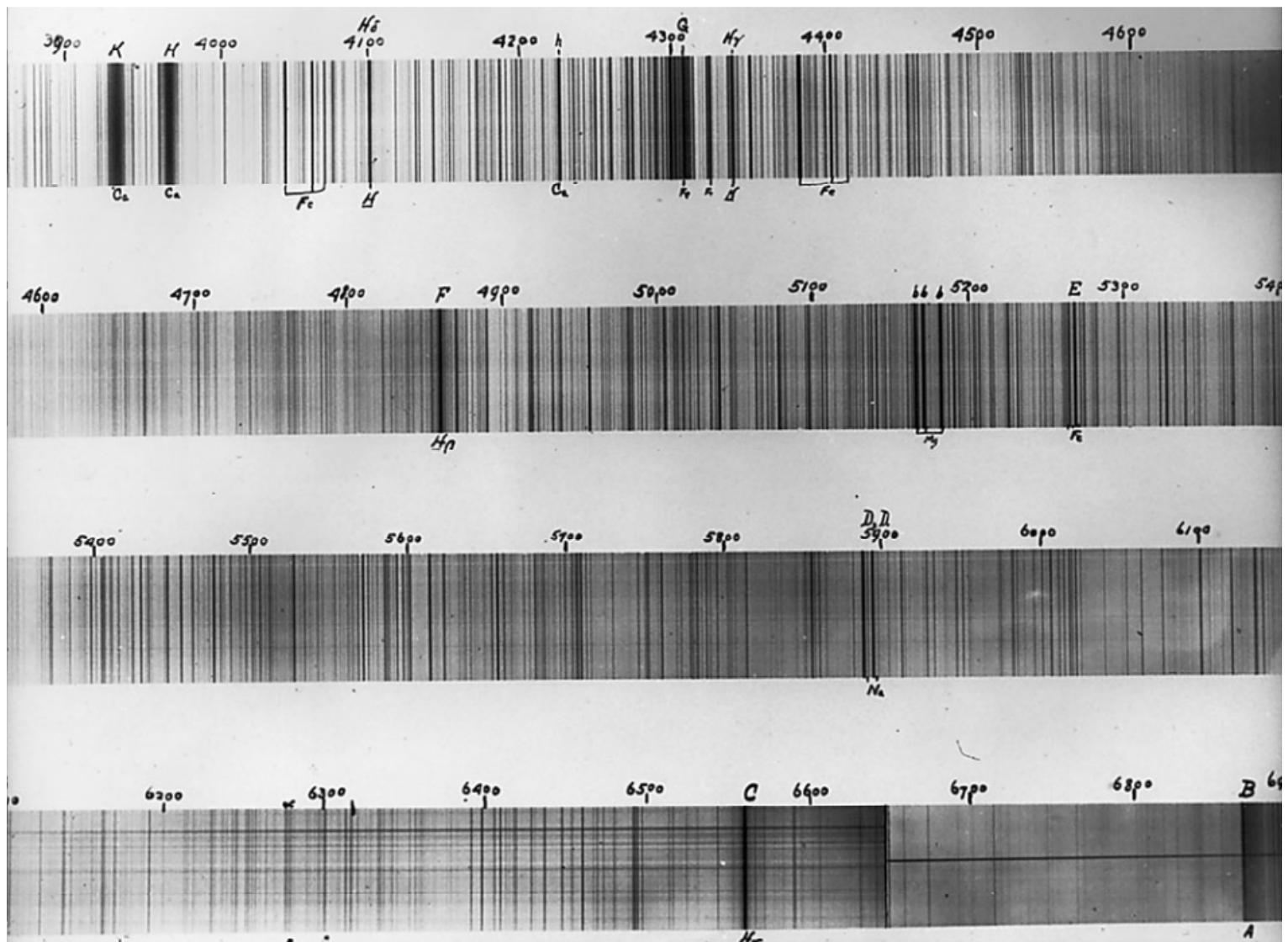
$$\frac{1}{\lambda} = R \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$$

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Bohr



Niels Henrik David Bohr
(1885-1962)



Niels Bohr Institute Group Photo 1960

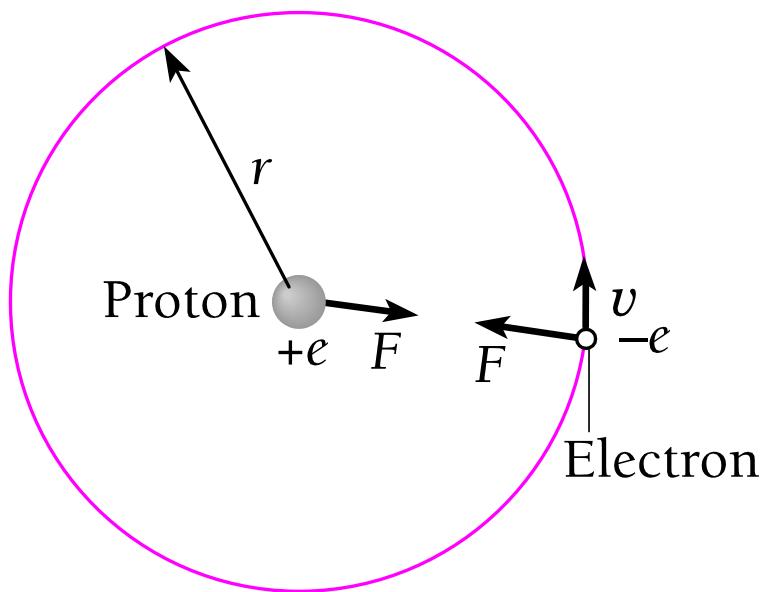


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Bohr의 원자 모델



$$\frac{mv^2}{r} = \frac{1}{4\pi\epsilon_0} \frac{e^2}{r^2}$$

$$v = \frac{e}{\sqrt{4\pi\epsilon_0 mr}}$$

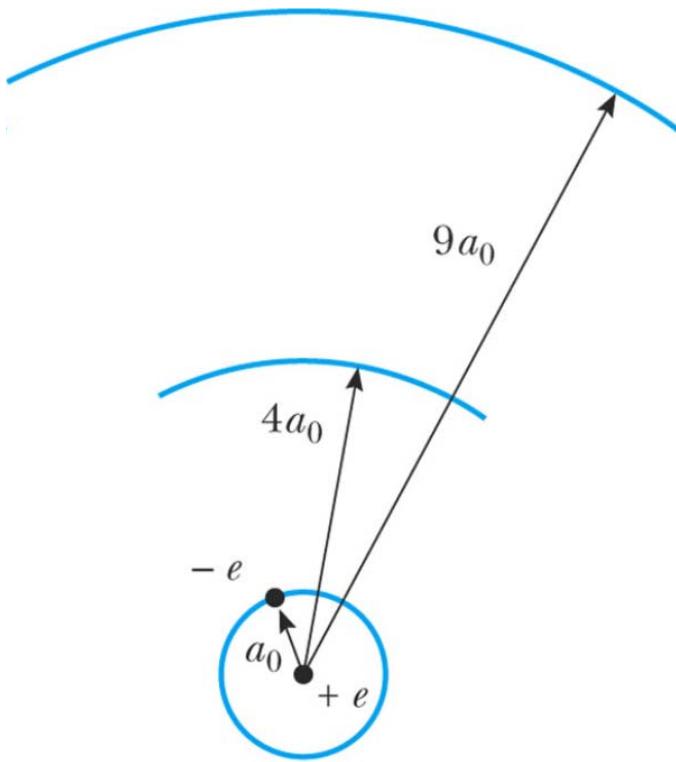
$$KE = \frac{1}{2}mv^2 \quad PE = -\frac{e^2}{4\pi\epsilon_0 r}$$

$$E = KE + PE = -\frac{e^2}{8\pi\epsilon_0 r}$$





Bohr의 원자 모델



$$r_n = n^2 a_0 \quad n = 1, 2, 3, \dots$$

$$mv r = n \hbar \quad (n = 1, 2, 3, \dots)$$

$$\left(\hbar = \frac{h}{2\pi} \right)$$

$$r_n = \frac{n^2 h^2 \epsilon_0}{\pi m e^2} = n^2 a_0 \quad (n = 1, 2, 3, \dots)$$

$$E_n = -\frac{me^4}{8\epsilon_0^2 h^2} \frac{1}{n^2} = -\frac{13.6 \text{eV}}{n^2} \quad (n = 1, 2, 3, \dots)$$

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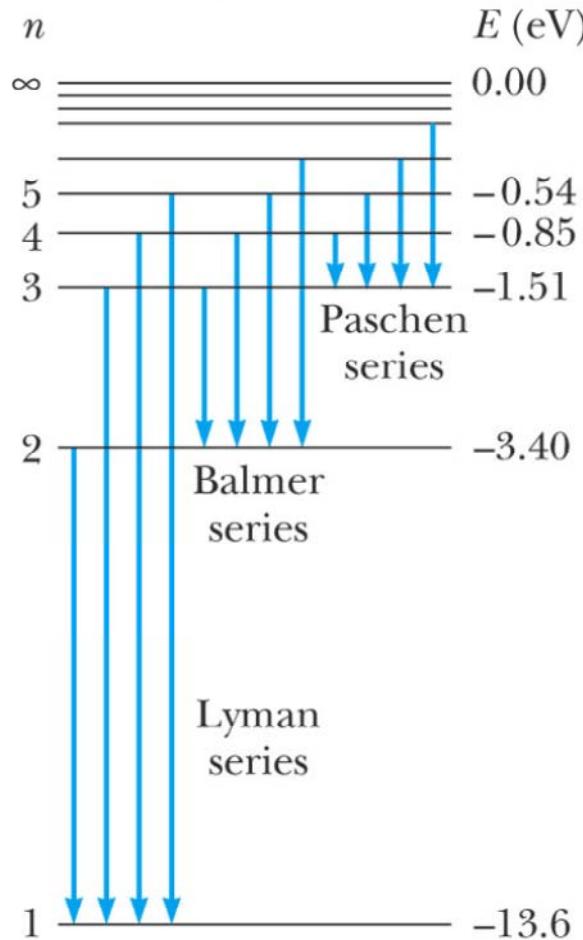


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Bohr의 원자 모델



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$$E_n = -\frac{me^4}{8\varepsilon_0^2 h^2} \frac{1}{n^2} = -\frac{13.6\text{eV}}{n^2} \quad (n=1,2,3,\dots)$$

$$E_i - E_f = hf$$

$$\frac{1}{\lambda} = \frac{f}{c} = \frac{me^4}{8\varepsilon_0^2 ch^3} \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$$



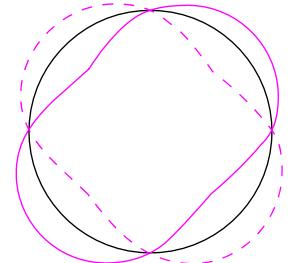
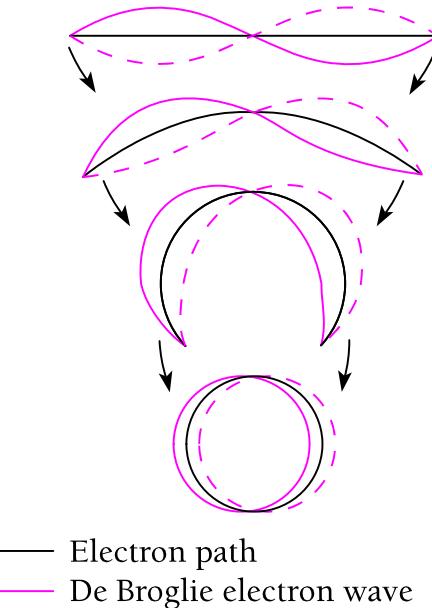


De Broglie

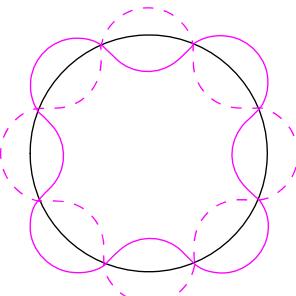
$$\lambda = \frac{h}{p}$$

$$2\pi r_n = n\lambda_n$$

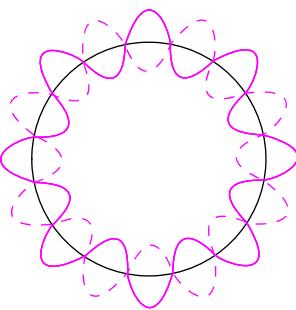
$$(n = 1, 2, 3, \dots)$$



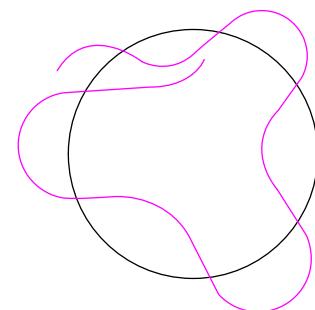
Circumference = 2 wavelengths



Circumference = 4 wavelengths



Circumference = 8 wavelengths



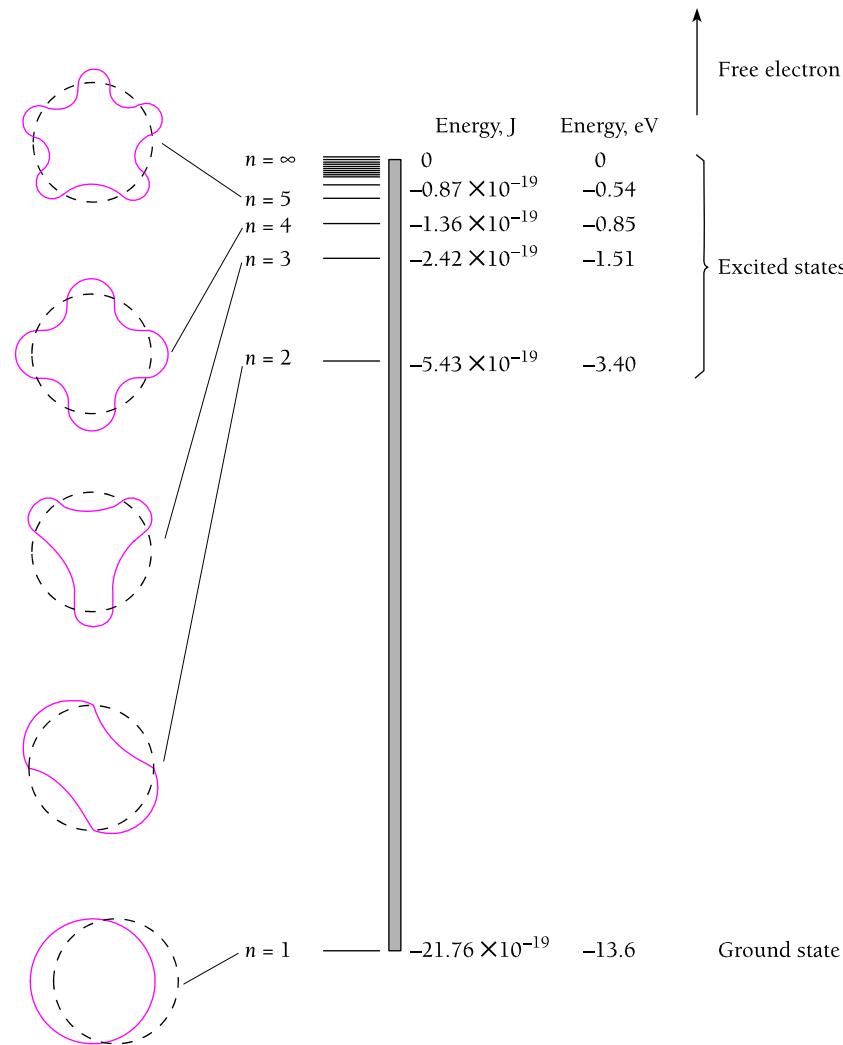


그림 4.15 수소 원자의 에너지 준위



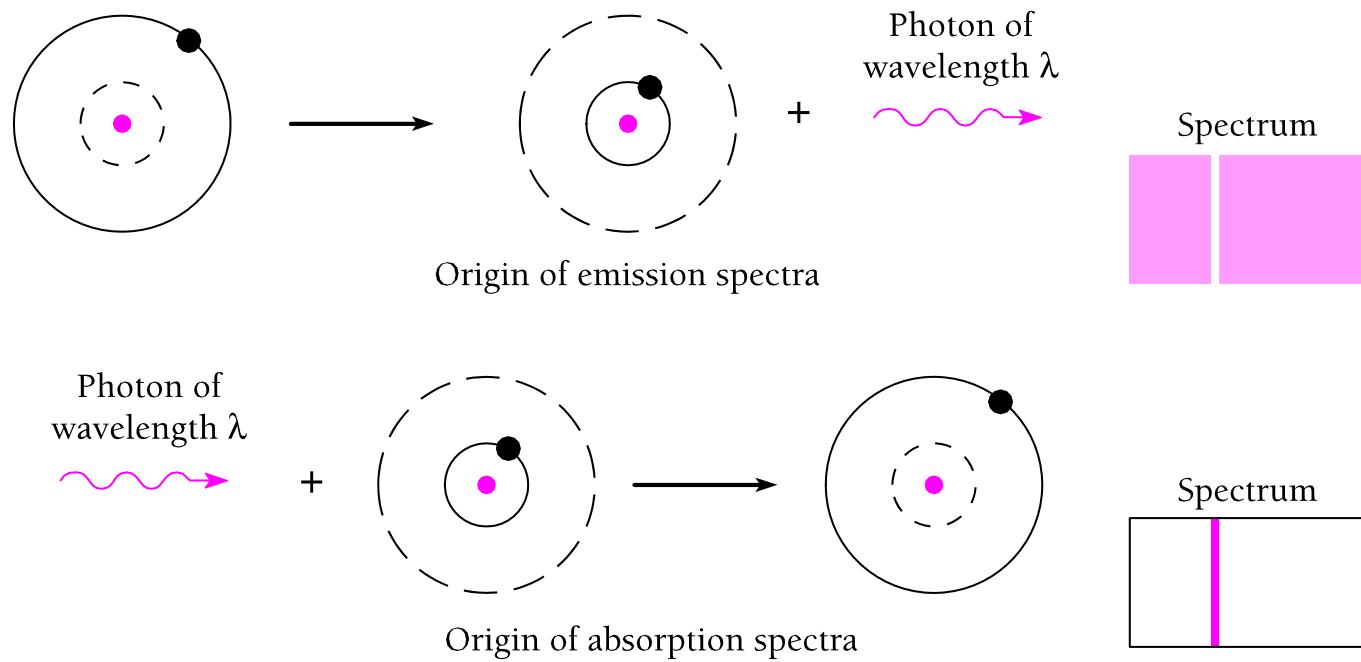


그림 4.19 방출 스펙트럼과 흡수 스펙트럼의 기원.

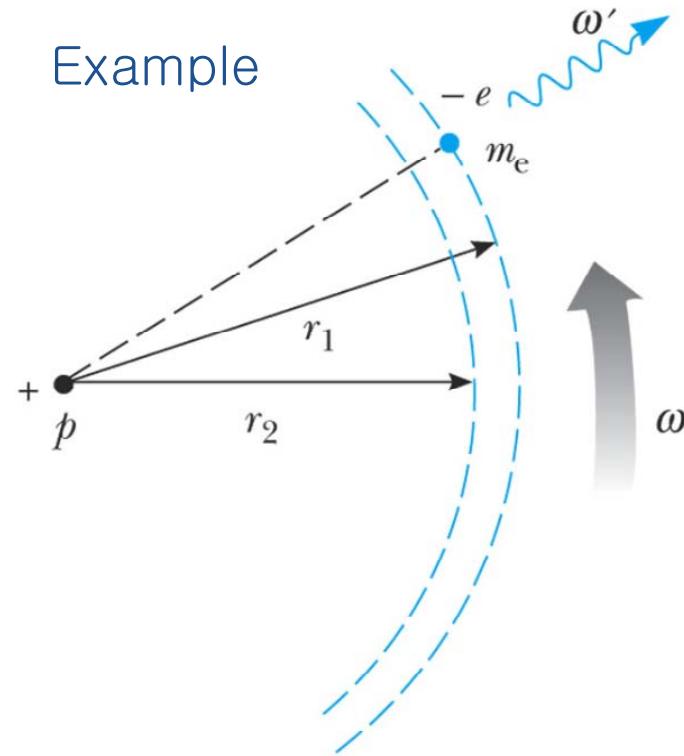




Correspondence Principle

$h \rightarrow 0$ or $n \rightarrow \infty$: Classical limit

Example



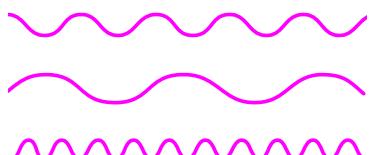
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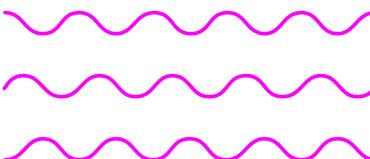
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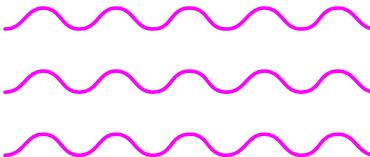
Laser의 특성



Ordinary light



Monochromatic,
incoherent light

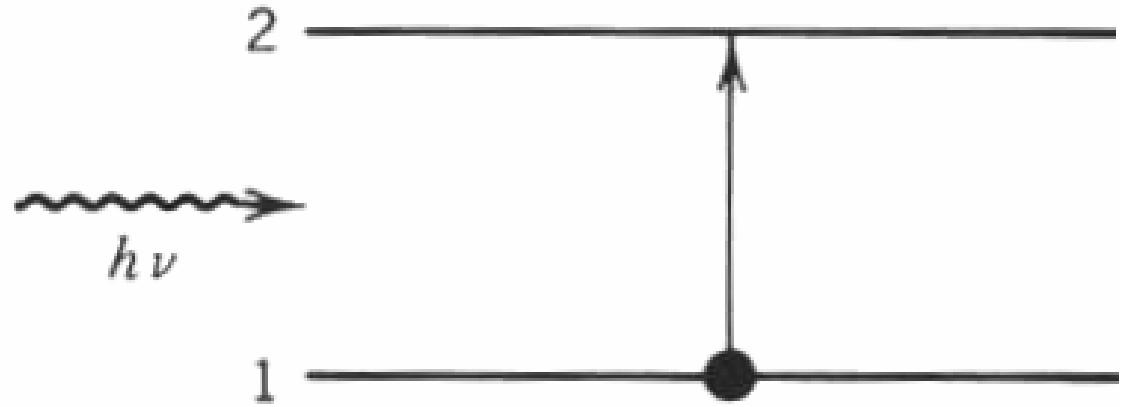


Monochromatic,
coherent light

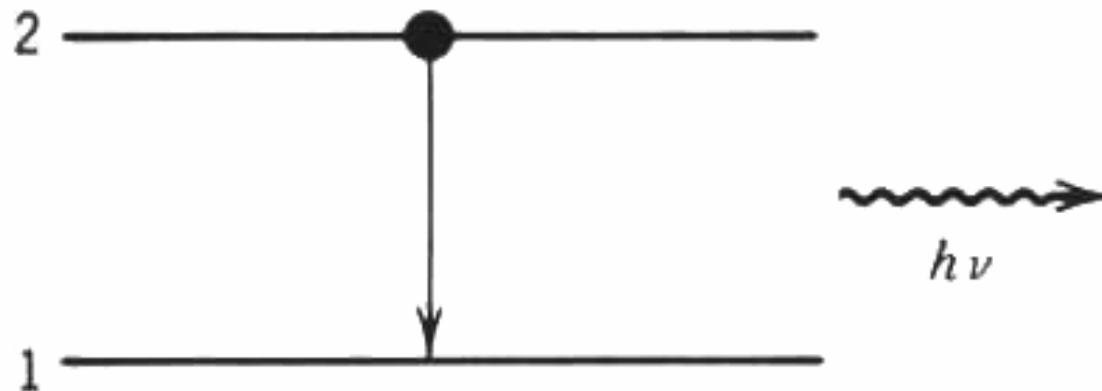
- 단색성
- 고휘도
- 지향성
- 가간섭성(Coherency)

그림 4.23 레이저는 모든 광파들이 같은 진동수를 가지고(단색광) 서로 위상이 같은(결맞음) 빛을 발생시킨다. 또한, 레이저빔은 매우 잘 평행하게 줄맞춤 된 빔이어서, 먼 거리를 지나가도 잘 퍼지지 않는다.





흡수(Absorption)

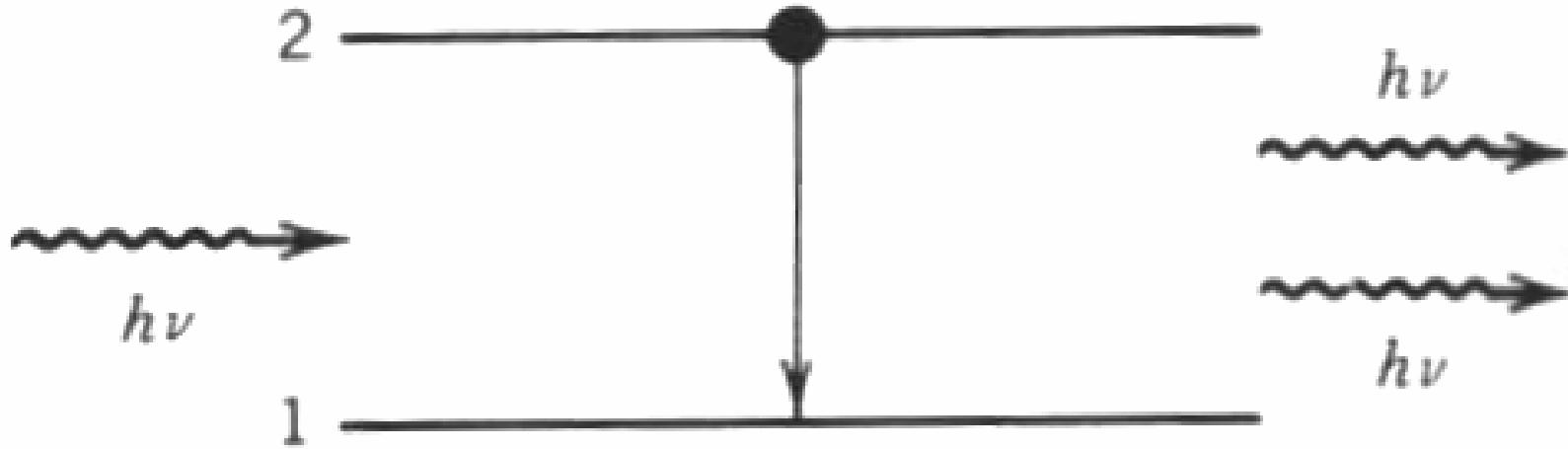


자발방출
(Spontaneous Emission)



유도방출(Stimulated Emission)

1917 Einstein



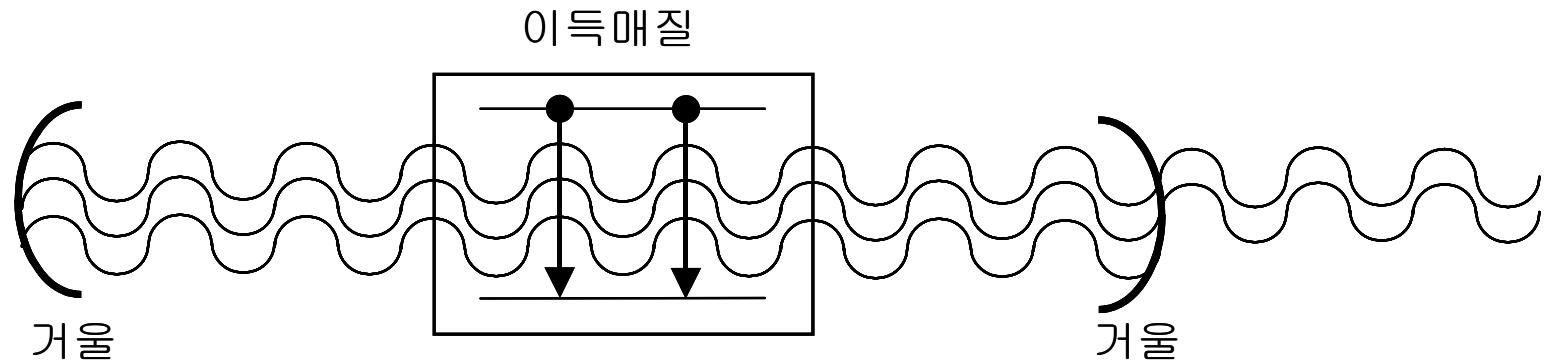
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Laser

Light Amplification by Stimulated Emission of Radiation





Laser의 역사

- 1917 Einstein – Stimulated Emission
- 1940s Fabrikant – 박사학위 논문
- 1950s Townes, Weber, Prokhorov, Basov – MASER
- 1954 Gordon, Townes, Zeiger – 최초의 maser
- 1957 Gould – LASER 특허출원
- 1958 Schawlow, Townes – “Infrared and Optical Masers” 논문
- 1960. 5. 16. Maiman – 최초의 레이저 동작 (Ruby laser)
- 1960. 12. 12. Javan, Bennett, Herriott – 최초의 cw 레이저 동작 (He-Ne laser)
- ... laser flood





Laser가 늦게 발명된 이유



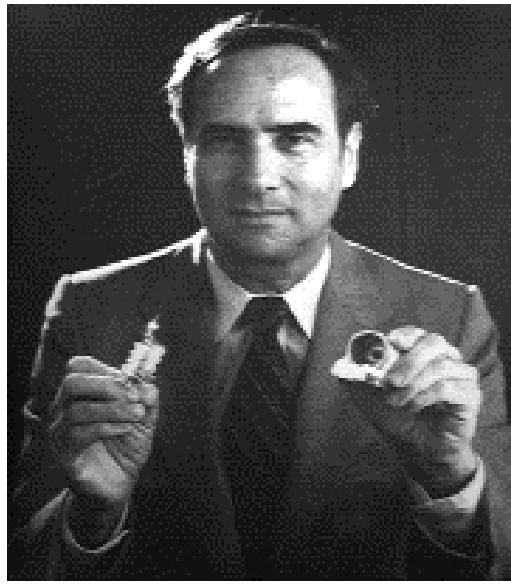
Charles H. Townes (1915-)

“대부분의 물리학자들은 전자공학과 증폭기에 대해서 몰랐었고, 전기공학자들은 대개 양자역학을 배우지 않았다. 하지만 제 2차 세계대전으로 인해 레이다 (radar) 개발을 위해 공학자들과 자연과학자들이 함께 일하게 되었고, 물리학자들이 전자공학에 접근할 수 있게 되었다.”





최초의 Laser – Ruby Laser



- ❖ 최초의 레이저 : 1960년, Maiman이 만든 세계 최초의 레이저,
루비를 레이저 매질로 사용

