

2018 Fall

“Phase Transformation *in* Materials”

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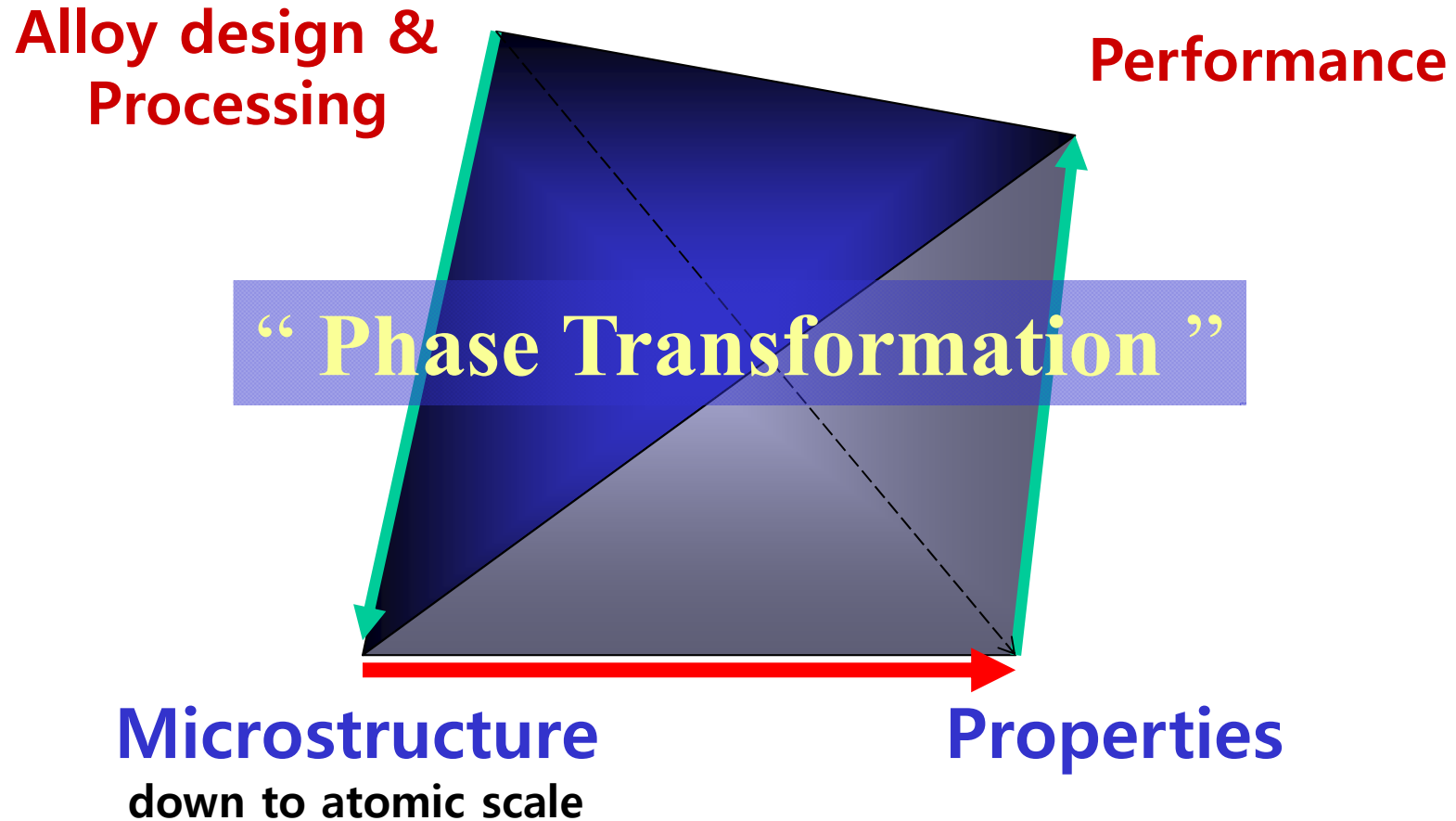
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Contents for previous class

Microstructure-Properties Relationships



“Tailor-made Materials Design”

Contents for previous class

Contents of this course_Phase transformation

**Background
to understand
phase
transformation**

(Ch1) Thermodynamics and Phase Diagrams

(Ch2) Diffusion: Kinetics

(Ch3) Crystal Interface and Microstructure

**Representative
Phase
transformation**

(Ch4) Solidification: Liquid \rightarrow Solid

(Ch5) Diffusional Transformations in Solid: Solid \rightarrow Solid

(Ch6) Diffusionless Transformations: Solid \rightarrow Solid

Contents for today's class I

Chapter 1

Thermodynamics and Phase Diagrams

- **Equilibrium**
- **Single component system**
 - Gibbs Free Energy
as a Function of Temp. and Pressure
- **Classification of phase transition**
- **Driving force for solidification**

Q1: “thermodynamic equilibrium”?

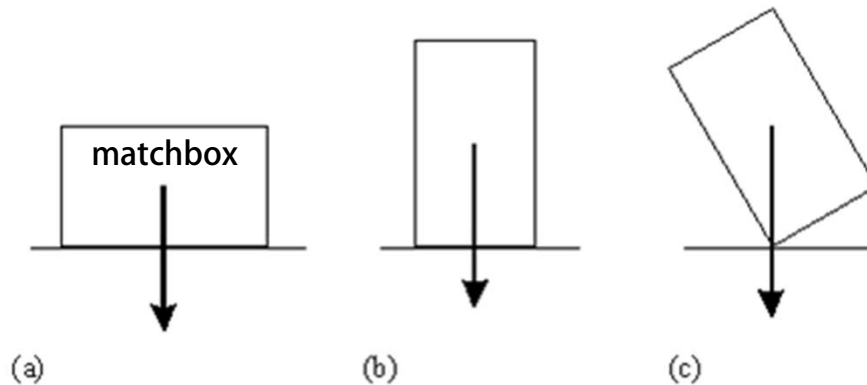
Lowest possible value of Gibb’s Free Energy

→ **The main use of thermodynamics in metallurgy is to allow the prediction of whether an alloy is in equilibrium.**

Chapter 1

Equilibrium

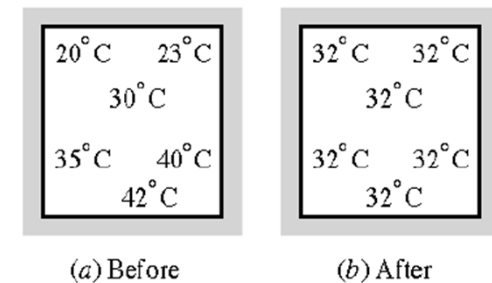
Mechanical equilibrium



: total potential energy of the system is a minimum.

Thermal equilibrium

: absence of temperature gradients in the system



Chemical equilibrium

: no further reaction occurs between the reacting substances
i.e. the forward and reverse rates of reaction are equal.

Thermodynamic equilibrium

: the system is under mechanical, thermal and chemical equilibrium

The properties of the system- P , T , V , concentrations-do not change with time.

Chapter 1.1

Relative Stability of a System \Rightarrow Gibbs Free Energy

$$\text{Gibbs free energy : } G = E + PV - TS = H - TS$$

Useful when P is constrained during thermodynamic process.

H : Enthalpy ; Measure of the heat content of the system

Mixture of one or more phases

$$H = E + PV$$

$$H \cong E \text{ for Condensed System}$$

E : Internal Energy, Kinetic + Potential Energy of a atom within the system

Kinetic Energy :

Atomic Vibration (Solid, Liquid)

Translational and Rotational Energy in liquid and gas.

Potential Energy : Interactions or Bonds between the atoms within the system

T : The Absolute Temperature

S : Entropy, The *Randomness* of the System

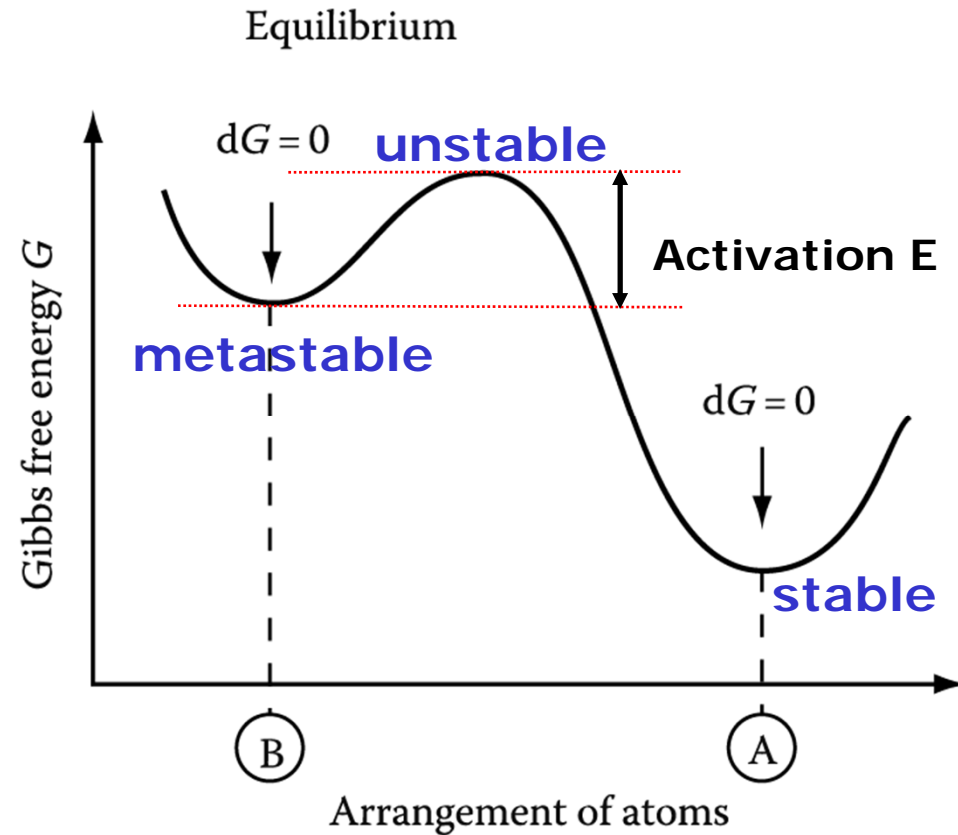
Chapter 1.1

Equilibrium

$$dG = 0$$

Lowest possible value of Gibb's Free Energy

No desire to change ad infinitum



Phase Transformation

$$\Delta G = G_2 - G_1 < 0$$

Q2: What is single component system?

Different elements or chemical compounds

1.2 Single component system

One element (Al, Fe)

One type of molecule (H₂O)

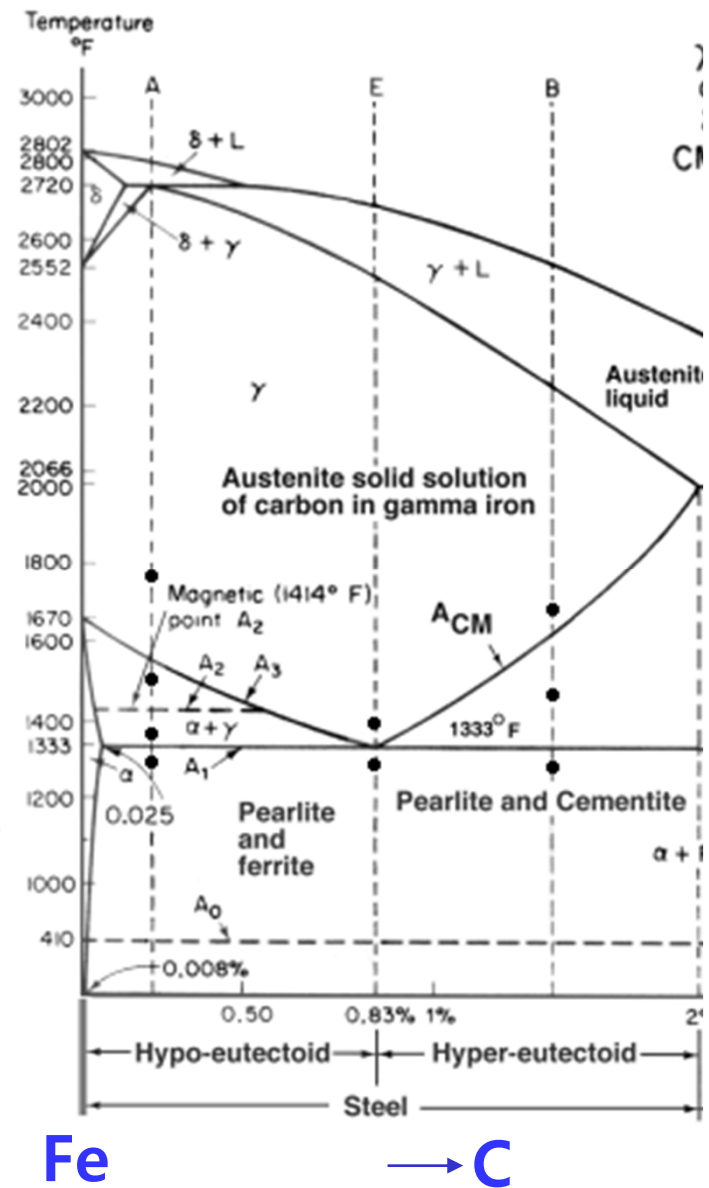
- Allotropic forms?

- How is phase stability measured?

$$G = H - TS$$



1.2.1 Gibbs Free Energy as a Function of Temp. 10



* What is the role of temperature on equilibrium?

Q3: C_v vs. C_p ?

Specific heat

(the quantity of heat (in joules) required to raise the temperature of substance by 1K)

at constant volume VS. at constant pressure

$$H = E + PV \longrightarrow dE = \delta Q - P \cdot dV$$

When V is constant,

$$\frac{\delta Q}{dT} = \frac{dE}{dT} + P \frac{dV}{dT} \xrightarrow{0} C_V = \left(\frac{\partial Q}{\partial T}\right)_V = \left(\frac{\partial E}{\partial T}\right)_V \quad C_V = \left(\frac{\partial E}{\partial T}\right)_V \quad \text{or} \quad E = \int C_V dT$$

Experimentally, it is easy to keep constant P (Cp) than constant V because it is difficult to make constant V (Cv) → pressure ex) 1 atm,

When pressure is constant,

$$H \equiv E + PV \longrightarrow dH = dE + PdV + VdP$$

$$= \delta Q - \delta w + PdV + VdP$$

$$= \delta Q - PdV + PdV + VdP$$

$$= \delta Q + VdP$$

$$dE = \delta Q - P \cdot dV$$

$$\rightarrow \frac{dH}{dT} = \frac{\delta Q}{dT} + V \frac{dP}{dT}$$

$$\rightarrow \frac{dP}{dT} = 0 \quad \text{when } P \text{ is constant}$$

$$\left(\frac{dH}{dT}\right)_P = \left(\frac{\delta Q}{dT}\right)_P = C_P$$

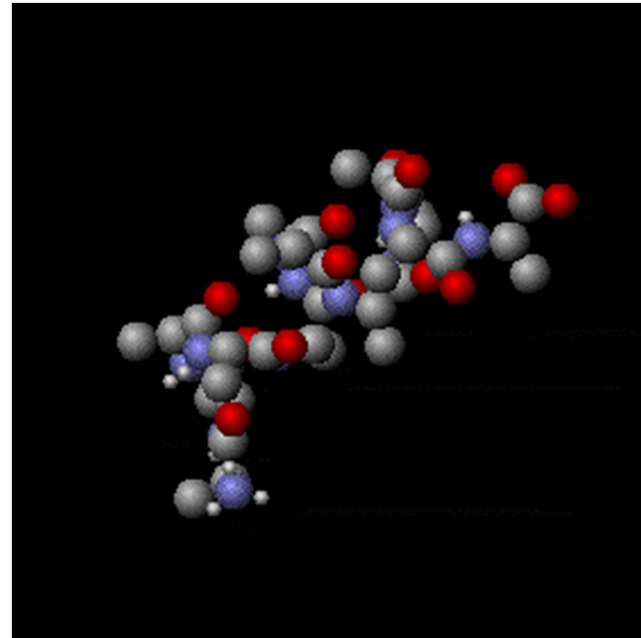
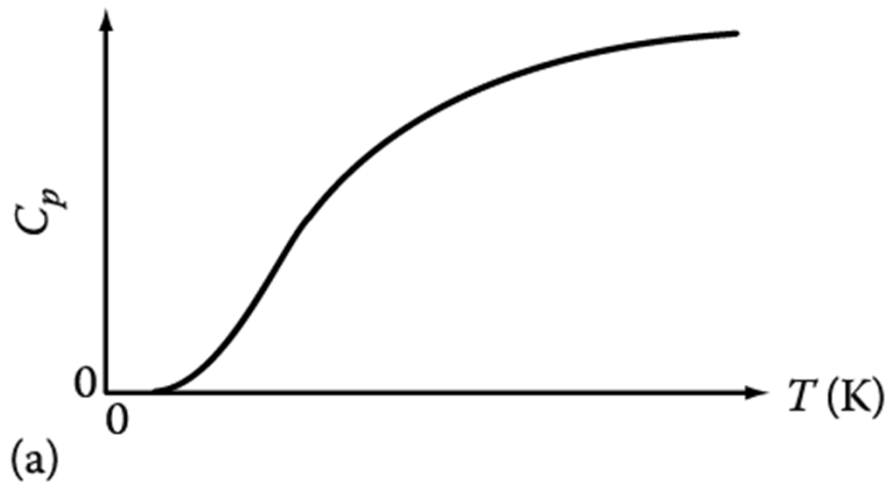
$$H = \int C_P dT$$

In alloy work one usually consider systems at constant pressure so that enthalpy changes are more important than energy changes.

C_p ; temperature-dependent function

$$C_p = a + bT + CT^{-2}$$

(empirical formula above room temp)



Molecules have internal structure because they are composed of atoms that have different ways of moving within molecules. Kinetic energy stored in these internal degrees of freedom contributes to a substance's specific heat capacity and not to its temperature.

*** What is the role of temperature on equilibrium?**

Q4: How is C_p related with H and S?

Draw the plots of (a) C_p vs. T , (b) H vs. T and (c) S vs. T .

How is C_p related with H and S ?

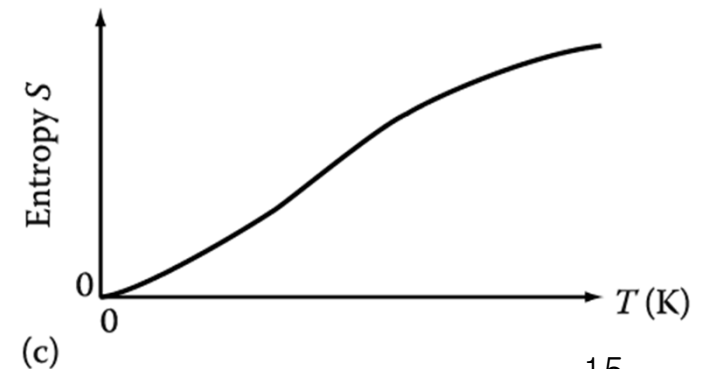
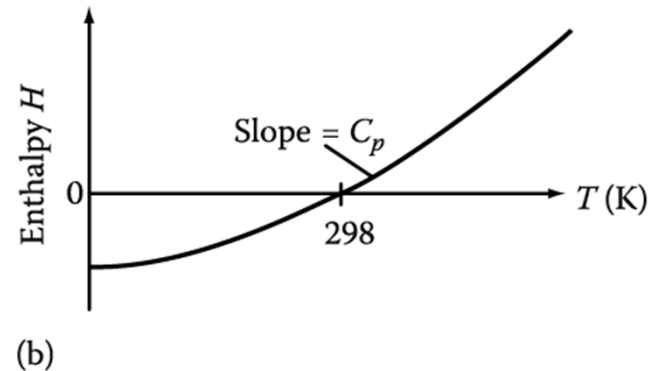
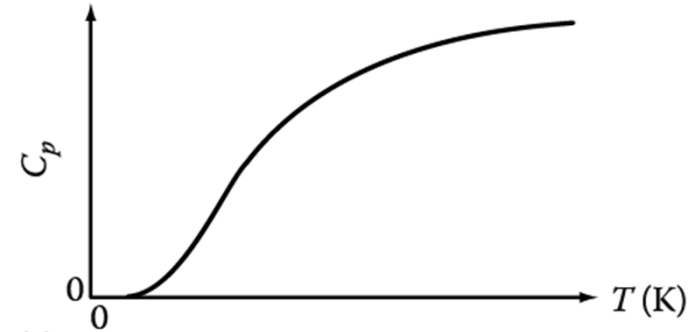
$$C_P = \left(\frac{\partial H}{\partial T} \right)_P \quad H = ? \quad H = \int_{298}^T C_P dT$$

$H = 0$ at 298K for a pure element
in its most stable state.

: When considering a phase transformation or a chemical reaction, the important thing is not the value of the thermodynamic function, but **the amount of change**.

Entropy : $S = \frac{q}{T}$

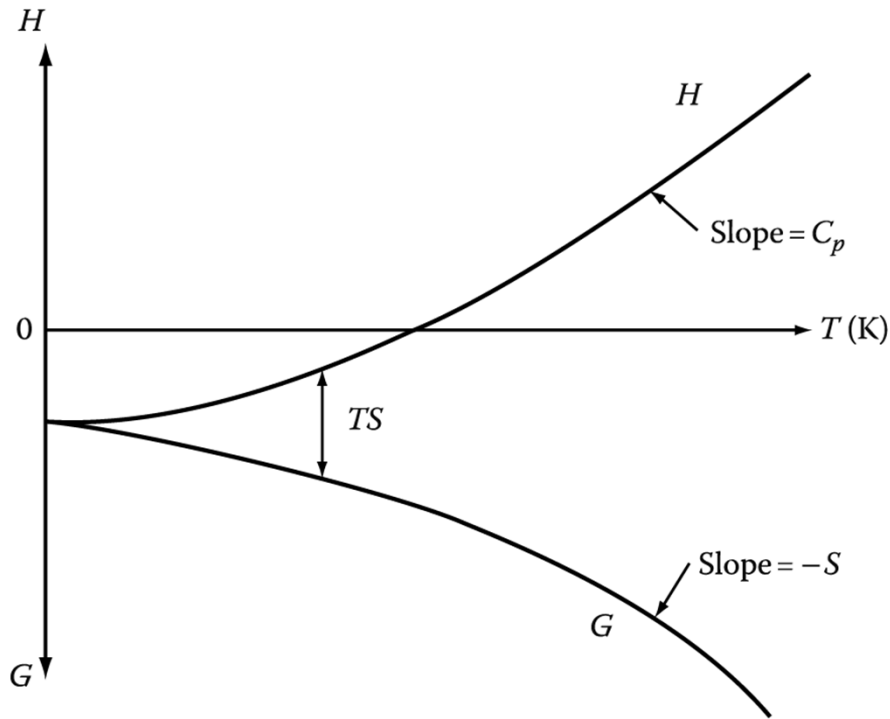
$$S = ? \quad \frac{C_P}{T} = \left(\frac{\partial S}{\partial T} \right)_P \quad S = \int_0^T \frac{C_P}{T} dT$$



*** What is the role of temperature on equilibrium?**

Q5: How to draw the plots of H vs.T and G vs. T in single component system?

Compare the plots of H vs.T and G vs. T.



$$G = G(T, P)$$

$$dG = \left(\frac{\partial G}{\partial T} \right)_P dT + \left(\frac{\partial G}{\partial P} \right)_T dP$$

$$G = H - TS$$

$$dG = dH - d(TS) = dE + d(PV) - d(TS)$$

$$dG = TdS - PdV + PdV + VdP - TdS - SdT$$

$$= VdP - SdT$$

$$\left(\frac{\partial G}{\partial T} \right)_P = -S, \quad \left(\frac{\partial G}{\partial P} \right)_T = V$$

$$dG = VdP - SdT$$

$$G(P, T) = G(P_0, T_0) + \int_{P_0}^{P_1} V(T_0, P) dP - \int_{T_0}^{T_1} S(P, T) dT$$

*** What is the role of temperature on equilibrium?**

Q6: G^S vs G^L as a function of temperature?

1.2.1 Gibbs Free Energy as a Function of Temp.

- Which is larger, H^L or H^S ?
- $H^L > H^S$ at all temp.
- Which is larger, S^L or S^S ?
- $S^L > S^S$ at all temp.

→ Gibbs free energy of the liquid decreases more rapidly with increasing temperature than that of the solid.

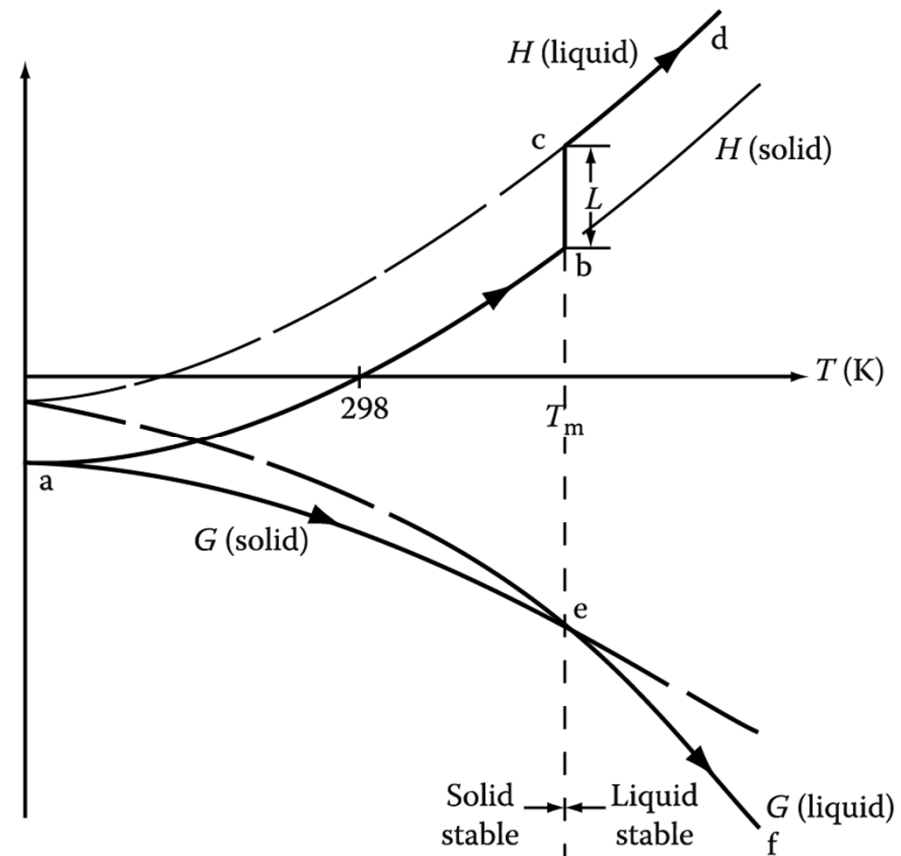


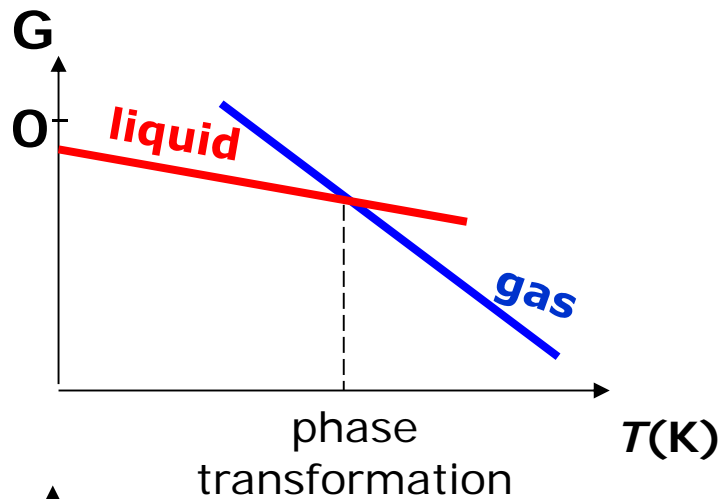
Fig. 1.4 Variation of enthalpy (H) and free energy (G) with temperature for the Solid and liquid phases of a pure metal. L is the latent heat of melting, T_m the Equilibrium melting temperature.

- Which is larger, G^L or G^S at low T?
- $G^L > G^S$ (at low Temp) and $G^S > G^L$ (at high Temp)

Considering P, T $G = G(T, P)$

$$dG = VdP - SdT$$

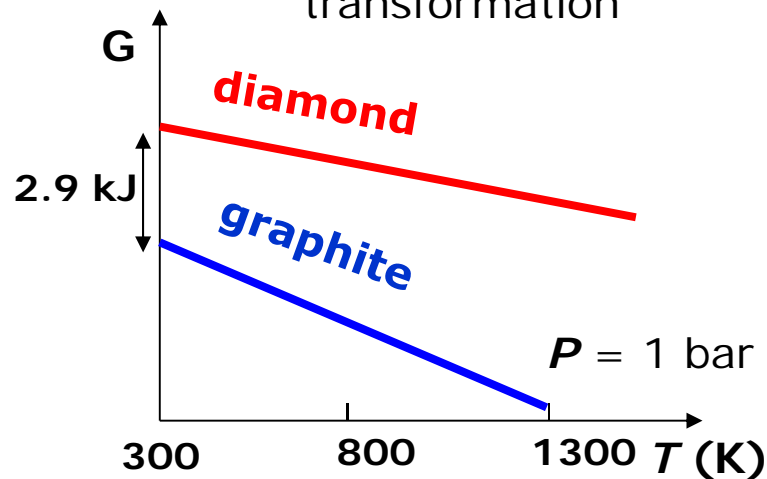
$$G(P, T) = G(P_0, T_0) + \int_{P_0}^{P_1} V(T_0, P)dP - \int_{T_0}^{T_1} S(P, T)dT$$



$$S(\text{water}) = 70 \text{ J/K}$$

$$S(\text{vapor}) = 189 \text{ J/K}$$

$$\left(\frac{\partial G}{\partial T} \right)_P = -S$$



$$S(\text{graphite}) = 5.74 \text{ J/K},$$

$$S(\text{diamond}) = 2.38 \text{ J/K},$$

Q7: What is the role of pressure on equilibrium?

$$* \textit{Clausius-Clapeyron Relation} : \left(\frac{dP}{dT} \right)_{eq} = \frac{\Delta H}{T_{eq} \Delta V}$$

(applies to all coexistence curves)

1.2.2 Pressure Effects

When the two phases with different molar volumes are in equilibrium, if the pressure changes, the equilibrium temperature T should also change with pressure.

If α & β phase are equilibrium,

$$dG^\alpha = V^\alpha dP - S^\alpha dT$$

$$dG^\beta = V^\beta dP - S^\beta dT$$

At equilibrium,

$$dG^\alpha = dG^\beta$$

$$\left(\frac{dP}{dT}\right)_{eq} = \frac{S^\beta - S^\alpha}{V^\beta - V^\alpha} = \frac{\Delta S}{\Delta V}$$

Here, $\Delta S = \frac{\Delta H}{T_{eq}}$

* *Clausius-Clapeyron Relation* :
$$\left(\frac{dP}{dT}\right)_{eq} = \frac{\Delta H}{T_{eq} \Delta V}$$

(applies to all coexistence curves)

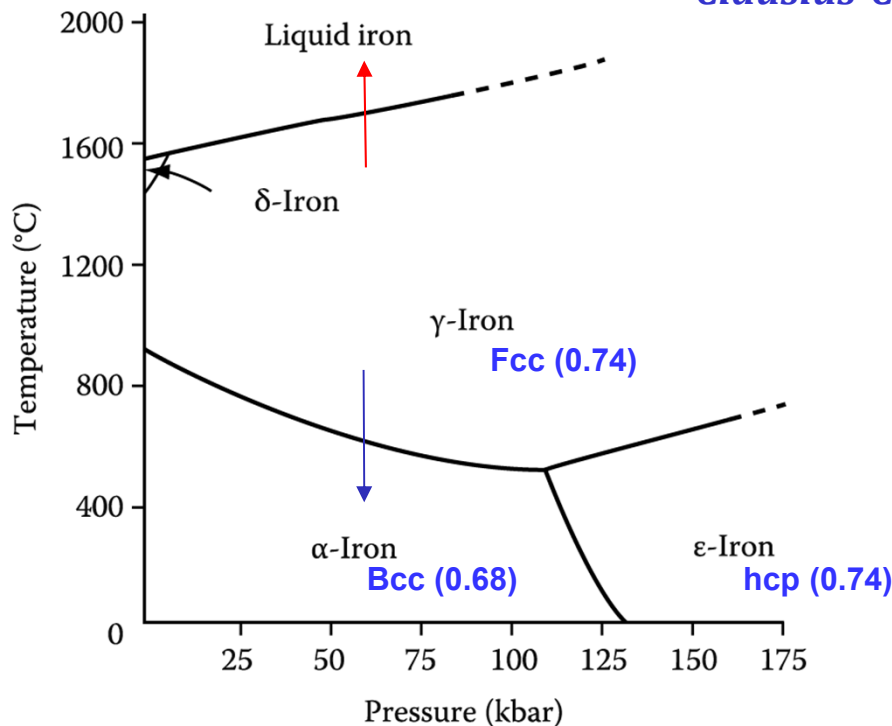


Fig. 1.5 Effect of pressure on the equilibrium phase diagram for pure iron

For, $\gamma \rightarrow$ liquid; $\Delta V (+)$, $\Delta H(+)$

$$\left(\frac{dP}{dT}\right) = \frac{\Delta H}{T_{eq} \Delta V} > 0$$

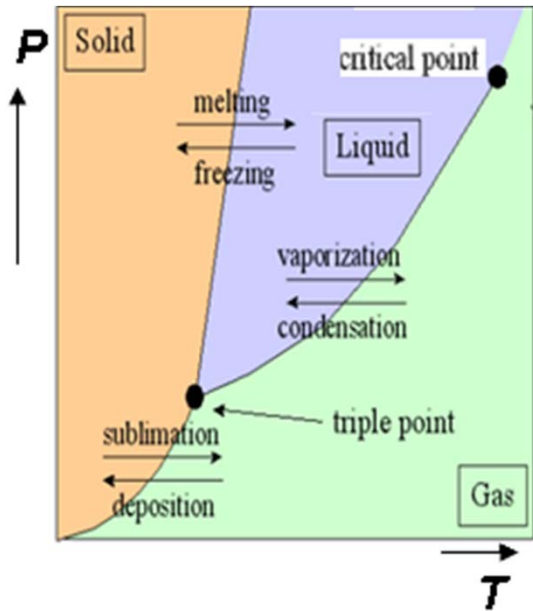
For, $\alpha \rightarrow \gamma$; $\Delta V (-)$, $\Delta H(+)$

$$\left(\frac{dP}{dT}\right) = \frac{\Delta H}{T_{eq} \Delta V} < 0$$

$$\left(\frac{\partial G}{\partial P}\right)_T = V$$

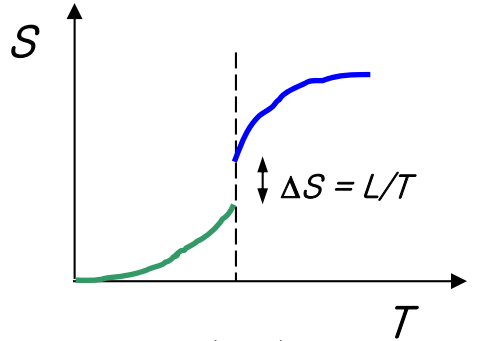
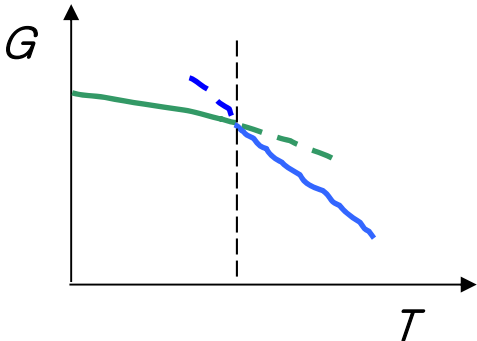
Q8: How to classify phase transition?

“First order transition” vs “Second order transition”

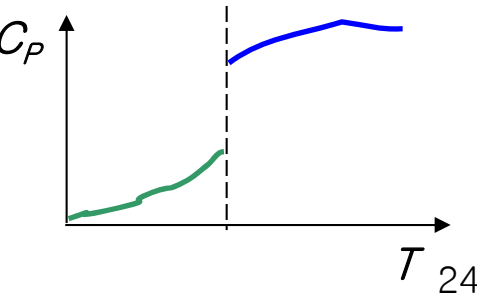


The First-Order Transition

Latent heat
Energy barrier
Discontinuous entropy, heat capacity



$$C_P = T \left(\frac{\partial S}{\partial T} \right)_{P,N}$$



• First Order Phase Transition at T_T :

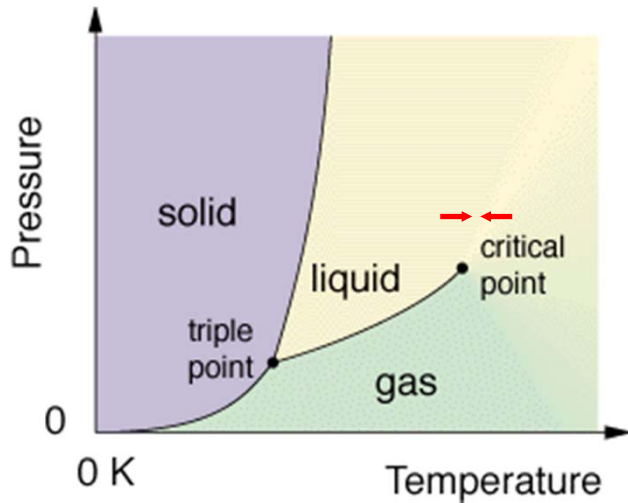
- G is **continuous** at T_T
- First derivatives of G (V, S, H) are **discontinuous** at T_T

$$V = \left(\frac{\partial G}{\partial P} \right)_T \quad S = - \left(\frac{\partial G}{\partial T} \right)_P \quad H = G - T \left(\frac{\partial G}{\partial T} \right)_P$$

- Second derivatives of G (α , β , C_p) are **discontinuous** at T_T

$$C_P = \left(\frac{\partial H}{\partial T} \right)_P \quad \alpha = \frac{1}{V} \left(\frac{\partial V}{\partial T} \right)_P \quad \beta = \frac{-1}{V} \left(\frac{\partial V}{\partial P} \right)_T$$

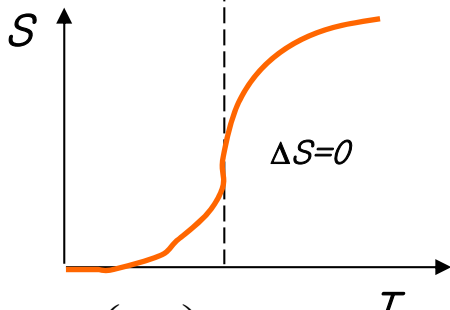
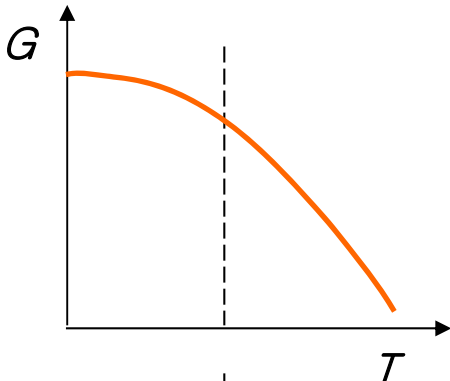
- Examples: Vaporization, Condensation, Fusion, Crystallization, Sublimation.



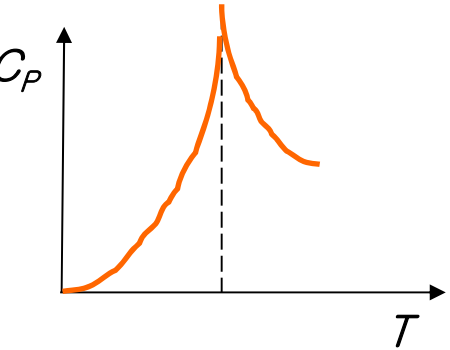
The Second Order Transition

No Latent heat
Continuous entropy

Second-order transition



$$C_p = T \left(\frac{\partial S}{\partial T} \right)_{P,N} \rightarrow \infty$$



• Second Order Phase Transition at T_T :

- G is continuous at T_T
- First derivatives of G (V, S, H) are continuous at T_T

$$V = \left(\frac{\partial G}{\partial P} \right)_T \quad S = - \left(\frac{\partial G}{\partial T} \right)_P \quad H = G - T \left(\frac{\partial G}{\partial T} \right)_P$$

- Second derivatives of G (α, β, C_p) are discontinuous at T_T

$$C_p = \left(\frac{\partial H}{\partial T} \right)_P \quad \alpha = \frac{1}{V} \left(\frac{\partial V}{\partial T} \right)_P \quad \beta = \frac{-1}{V} \left(\frac{\partial V}{\partial P} \right)_T$$

- **Examples:** Order-Disorder Transitions in Metal Alloys, Onset of Ferromagnetism, Ferroelectricity, Superconductivity.

**Q9: What is the driving force for
“Solidification: Liquid → Solid”?**

1.2.3 Driving force for solidification

$$G^L = H^L - TS^L$$

$$G^S = H^S - TS^S$$

$$\Delta G = \Delta H - T \Delta S$$

$$L : \Delta H = H^L - H^S$$

(Latent heat)

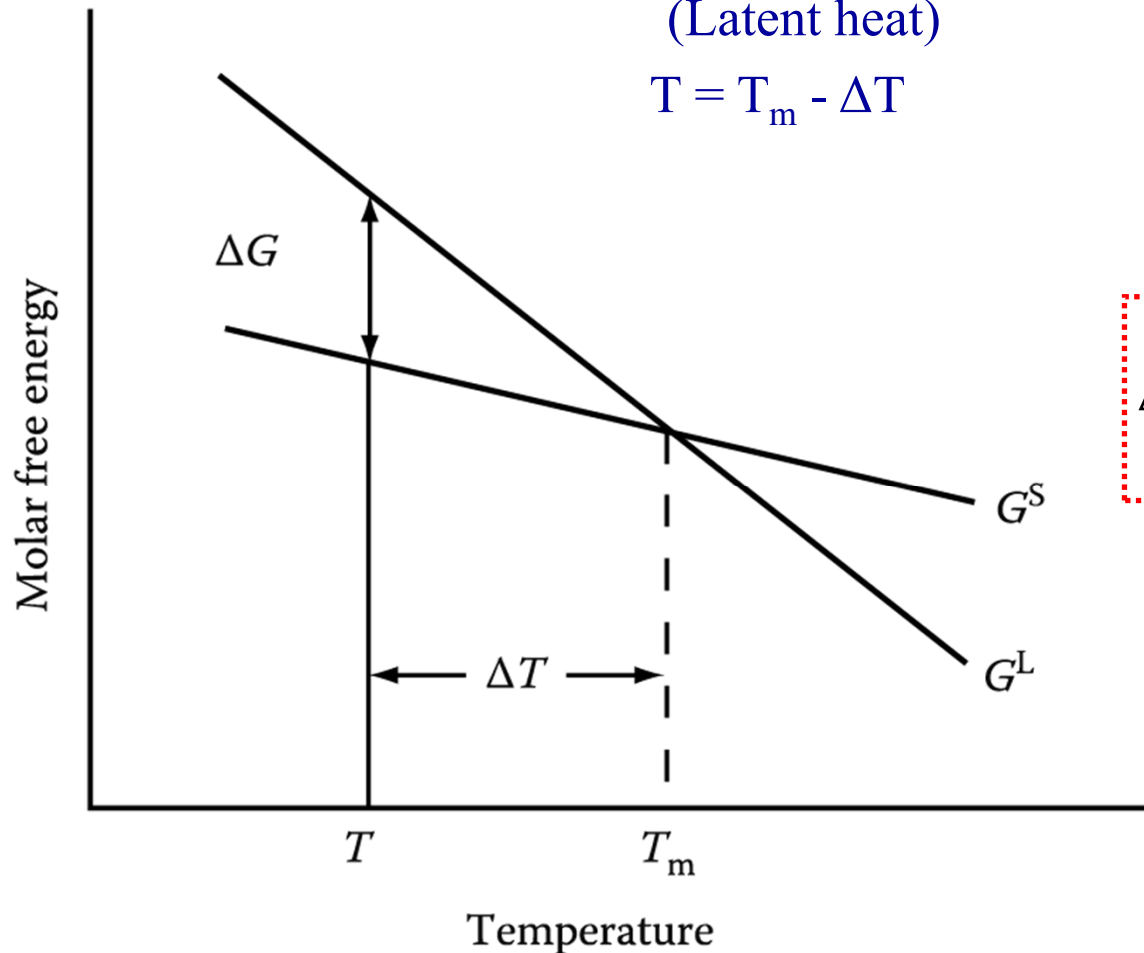
$$T = T_m - \Delta T$$

$$\Delta G = 0 = \Delta H - T_m \Delta S$$

$$\Delta S = \Delta H / T_m = L / T_m$$

$$\Delta G = L - T(L/T_m) \approx (L\Delta T) / T_m$$

(eq. 1.17)



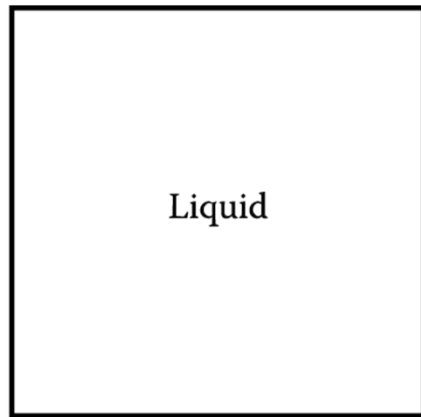
$$\Delta G = \frac{L\Delta T}{T_m}$$

4. Solidification: Liquid \longrightarrow Solid



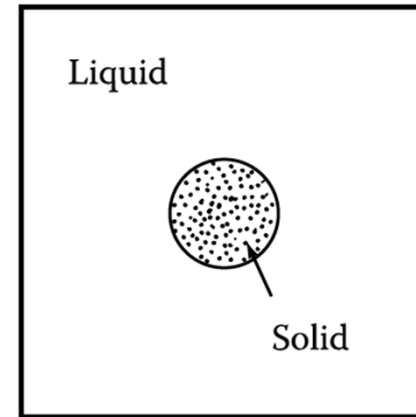
4 Fold Anisotropic Surface Energy/2 Fold Kinetics, Many Seeds

4.1.1. Homogeneous Nucleation



(a) G_1

$$G_1 = (V_S + V_L)G_V^L$$



(b) $G_2 = G_1 + \Delta G$

$$G_2 = V_S G_V^S + V_L G_V^L + A_{SL} \gamma_{SL}$$

G_V^S, G_V^L : free energies per unit volume

$$\Delta G = G_2 - G_1 = -V_S (G_V^L - G_V^S) + A_{SL} \gamma_{SL}$$

for spherical nuclei (isotropic) of radius : r

$$\Delta G_r = -\frac{4}{3} \pi r^3 \Delta G_V + 4\pi r^2 \gamma_{SL}$$

4.1.1. Homogeneous Nucleation

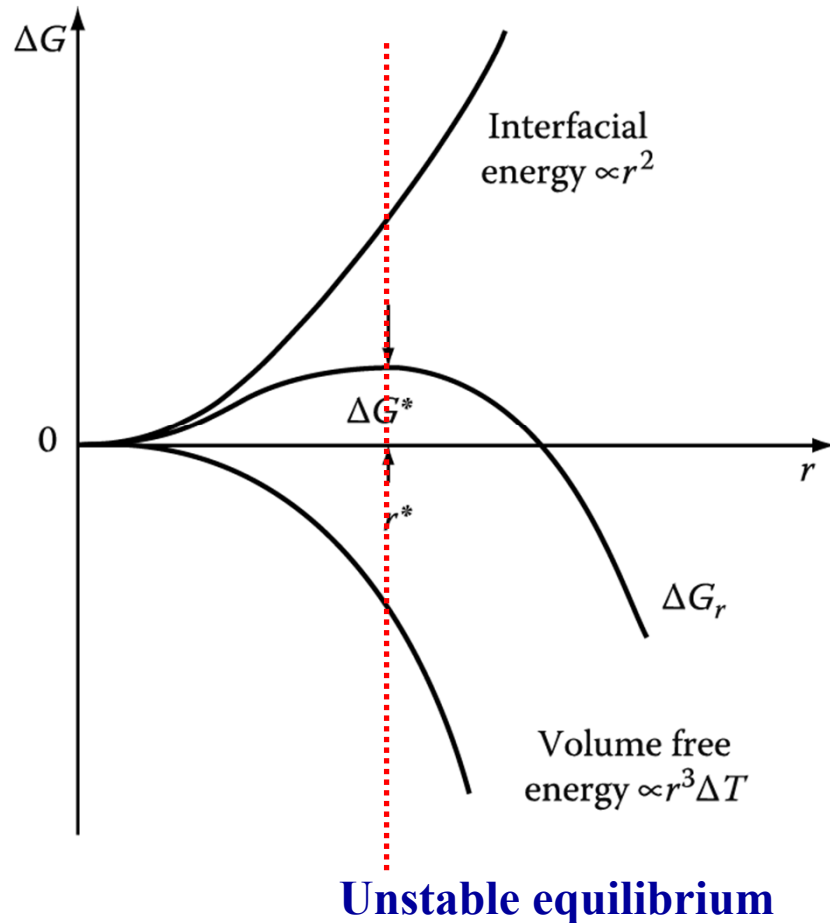


Fig. 4.2 The free energy change associated with homogeneous nucleation of a sphere of radius r .

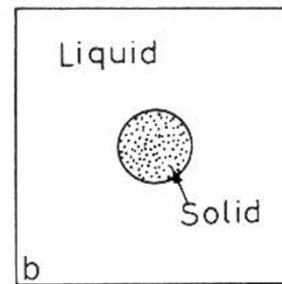
Why r^* is not defined by $\Delta G_r = 0$?

$r < r^*$: unstable (lower free E by reduce size)

$r > r^*$: stable (lower free E by increase size)

r^* : critical nucleus size

$r^* \implies dG=0$

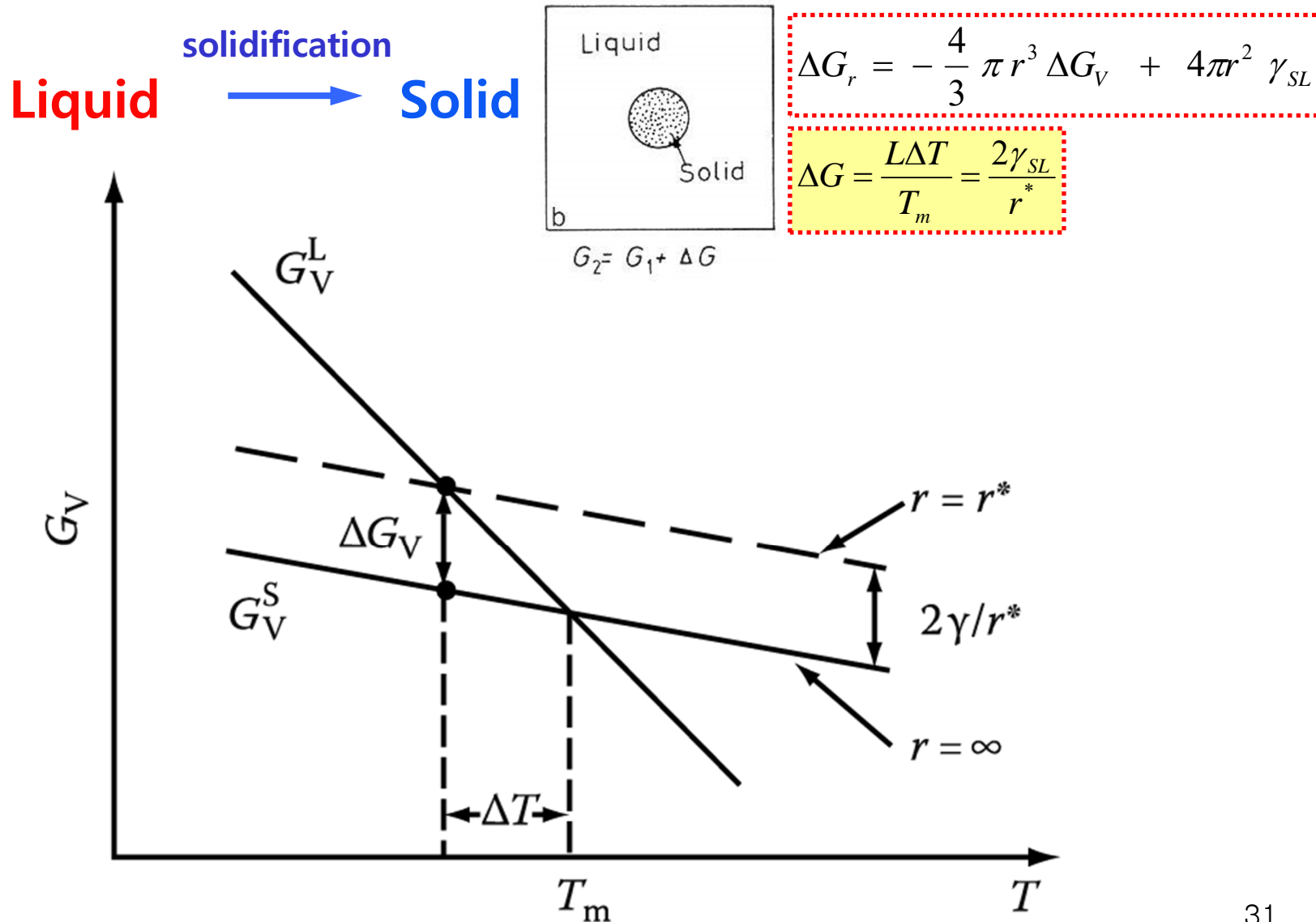


$$G_2 = G_1 + \Delta G$$

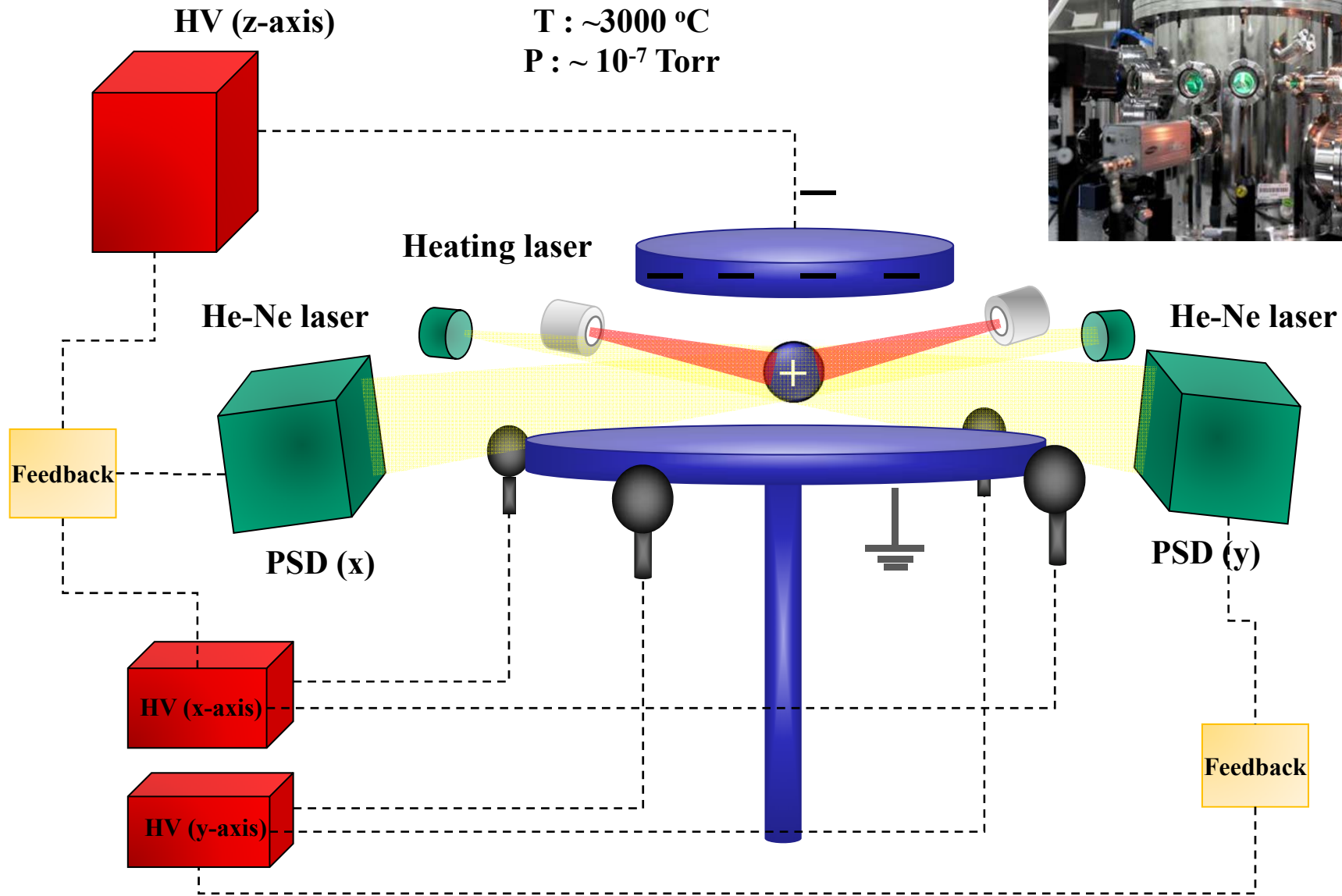
$$\Delta G_r = -\frac{4}{3} \pi r^3 \Delta G_V + 4\pi r^2 \gamma_{SL}$$

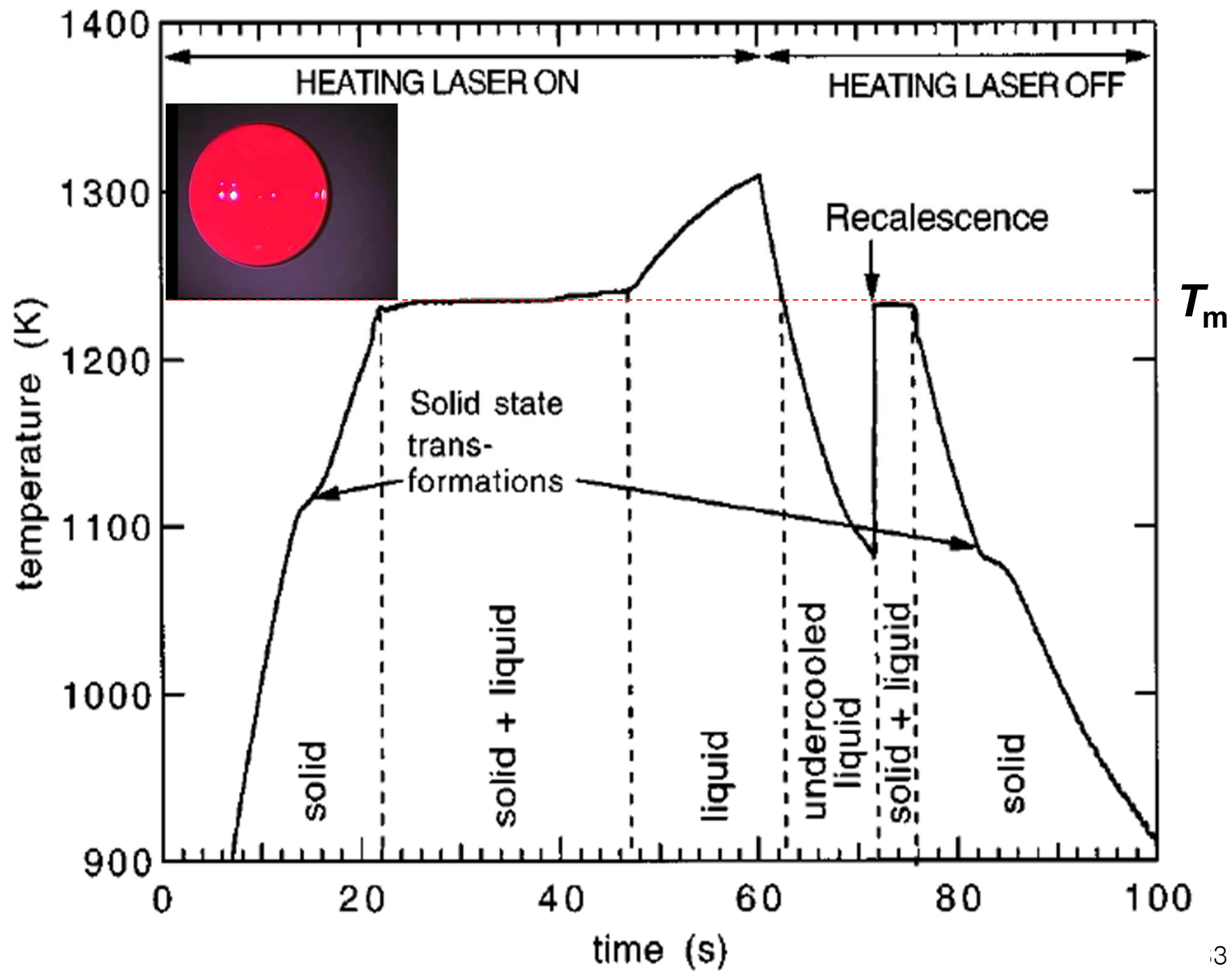
$$\Delta G^* = \frac{2\gamma_{SL}}{r^*}$$

1.2.3 Driving force for solidification

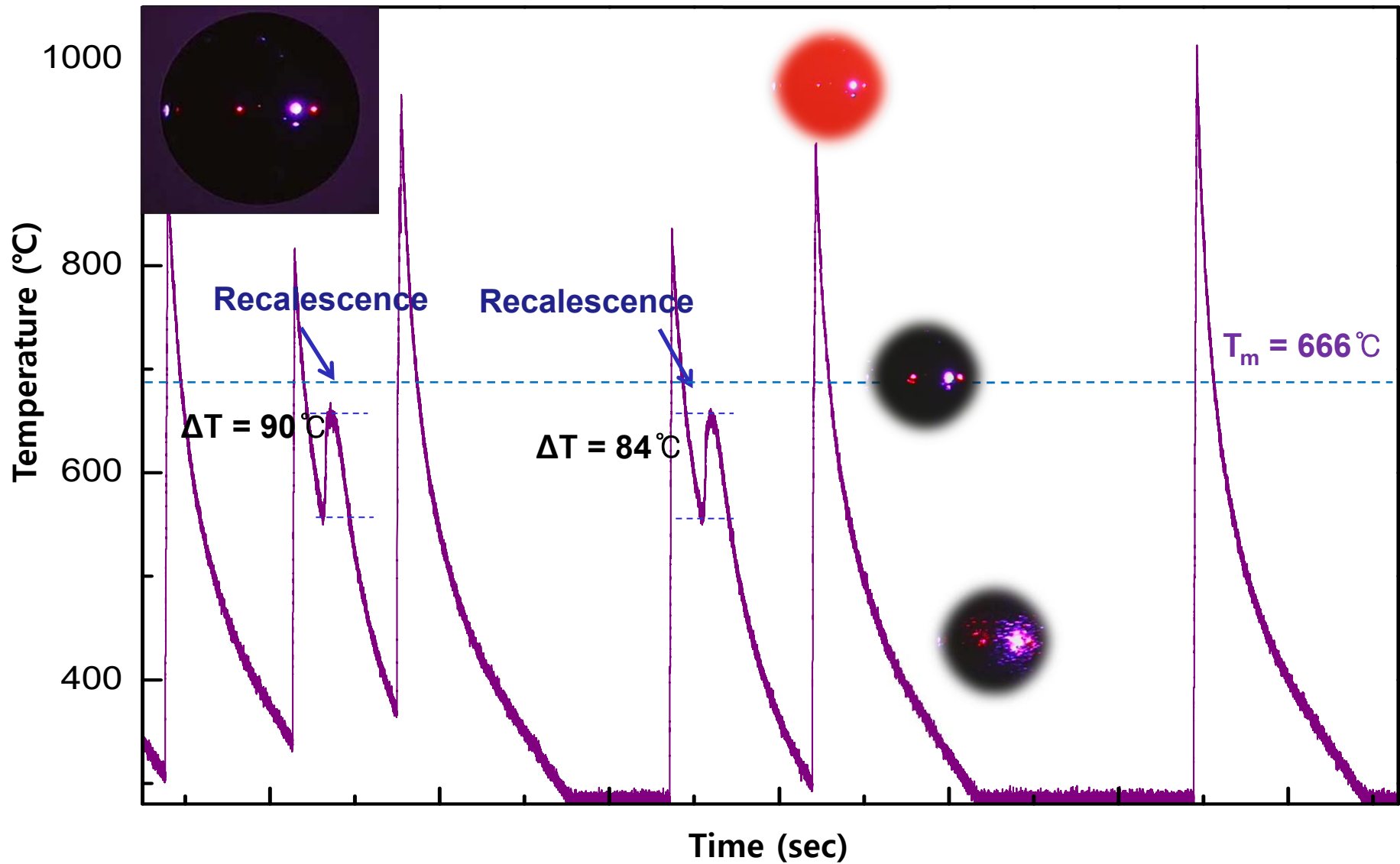


Electrostatic levitation in KRISS





Cyclic cooling curves of $\text{Zr}_{41.2}\text{Ti}_{13.8}\text{Cu}_{12.5}\text{Ni}_{10}\text{Be}_{22.5}$



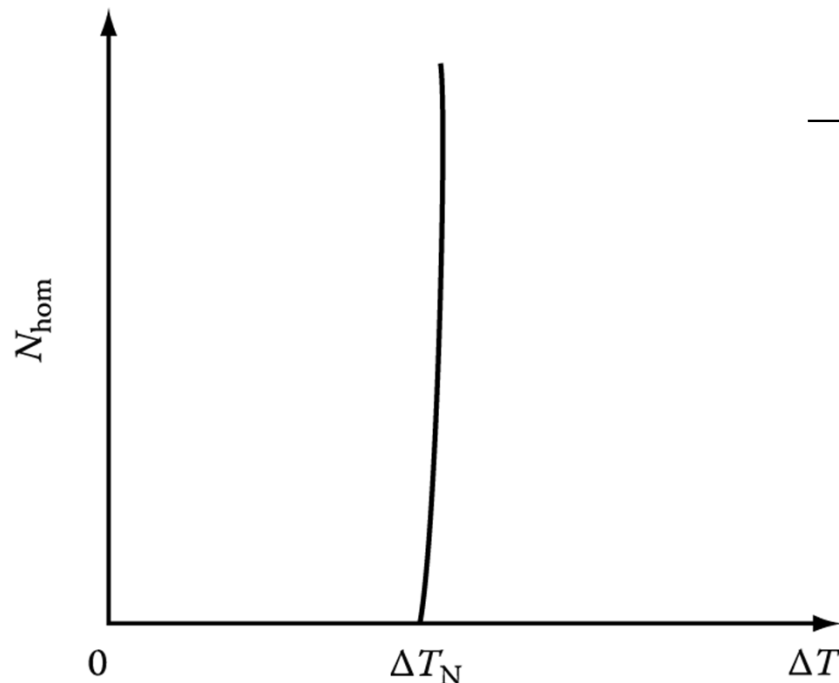
* The homogeneous nucleation rate - kinetics

$$I = \frac{nkT}{h} \exp\left(-\frac{G_A}{kT}\right) \exp\left(-\frac{16\pi\sigma_{SL}^3 T_E^2}{3L^2(\Delta T)^2 kT}\right)$$

: insensitive to Temp.

How do we define nucleation temperature, ΔT_N ?

$$N_{\text{hom}} \sim \frac{1}{\Delta T^2}$$



→ critical value for detectable nucleation

- critical supersaturation ratio
- critical driving force
- critical supercooling

→ for most metals, $\Delta T_N \sim 0.2 T_m$ (i.e. $\sim 200\text{K}$)

The homogeneous nucleation rate as a function of undercooling ΔT . ΔT_N is the **critical undercooling** for homogeneous nucleation.

* Relationship between Maximum Supercoolings and T_m

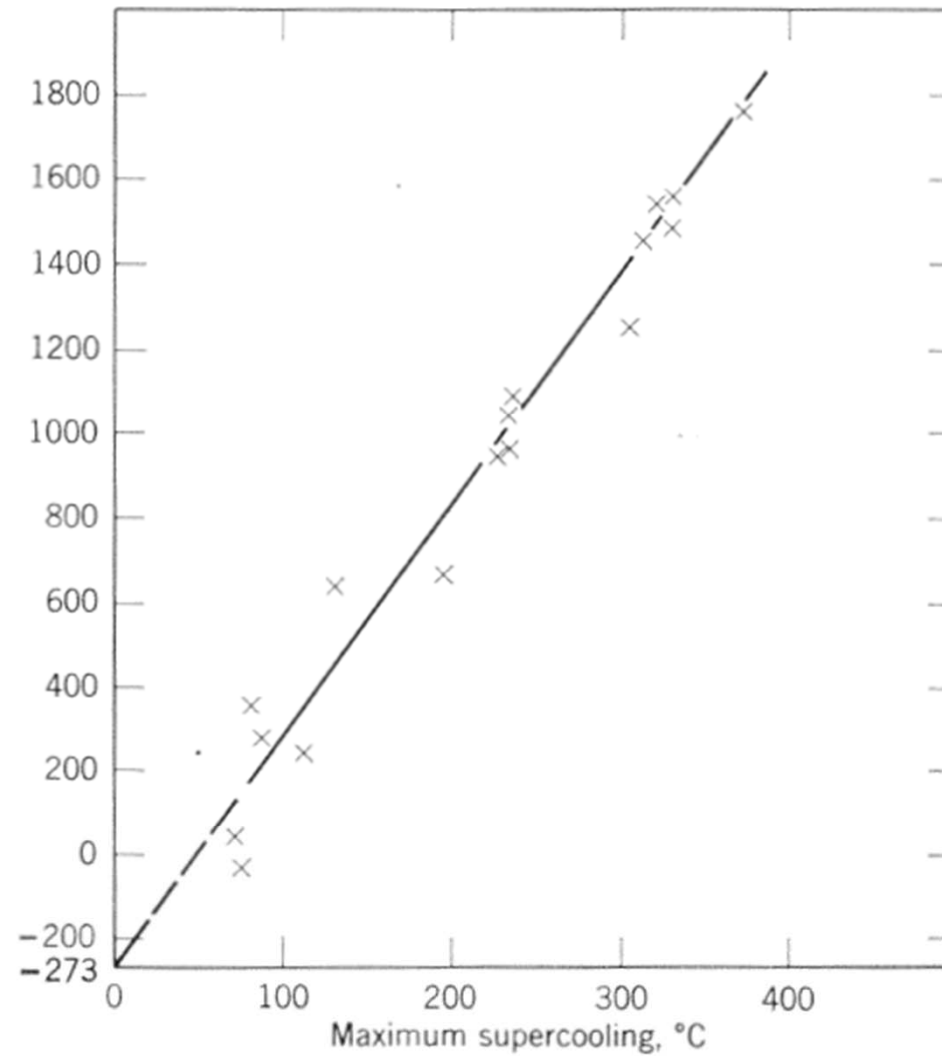


Fig. 3.7. Maximum supercooling as a function of melting point. (From *Thermodynamics in Physical Metallurgy*, American Society for Metals, Cleveland, 1911, p. 11.)

Solidification: Liquid \longrightarrow Solid

- casting & welding
- single crystal growth
- directional solidification
- rapid solidification

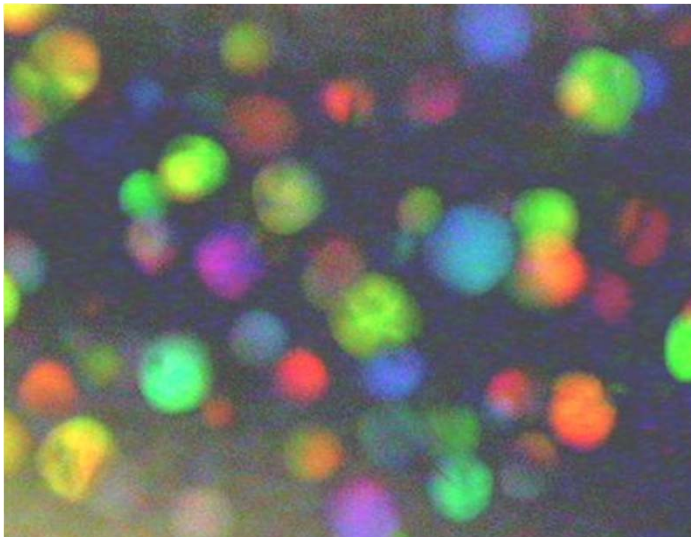
Nucleation in Pure Metals

$$T_m : G_L = G_S$$

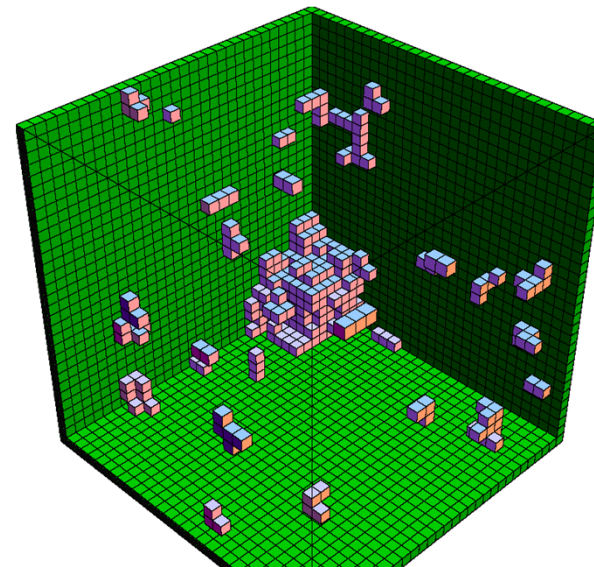
- Undercooling (supercooling) for nucleation: 250 K \sim 1 K

<Types of nucleation>

- Homogeneous nucleation



- Heterogeneous nucleation



**Q10: What is the driving force for
“Melting: Solid → Liquid”?**

* Driving force for melting

$$G^L = H^L - TS^L$$

$$G^S = H^S - TS^S$$

$$\Delta G = \Delta H - T \Delta S$$

$$L : \Delta H = H^L - H^S$$

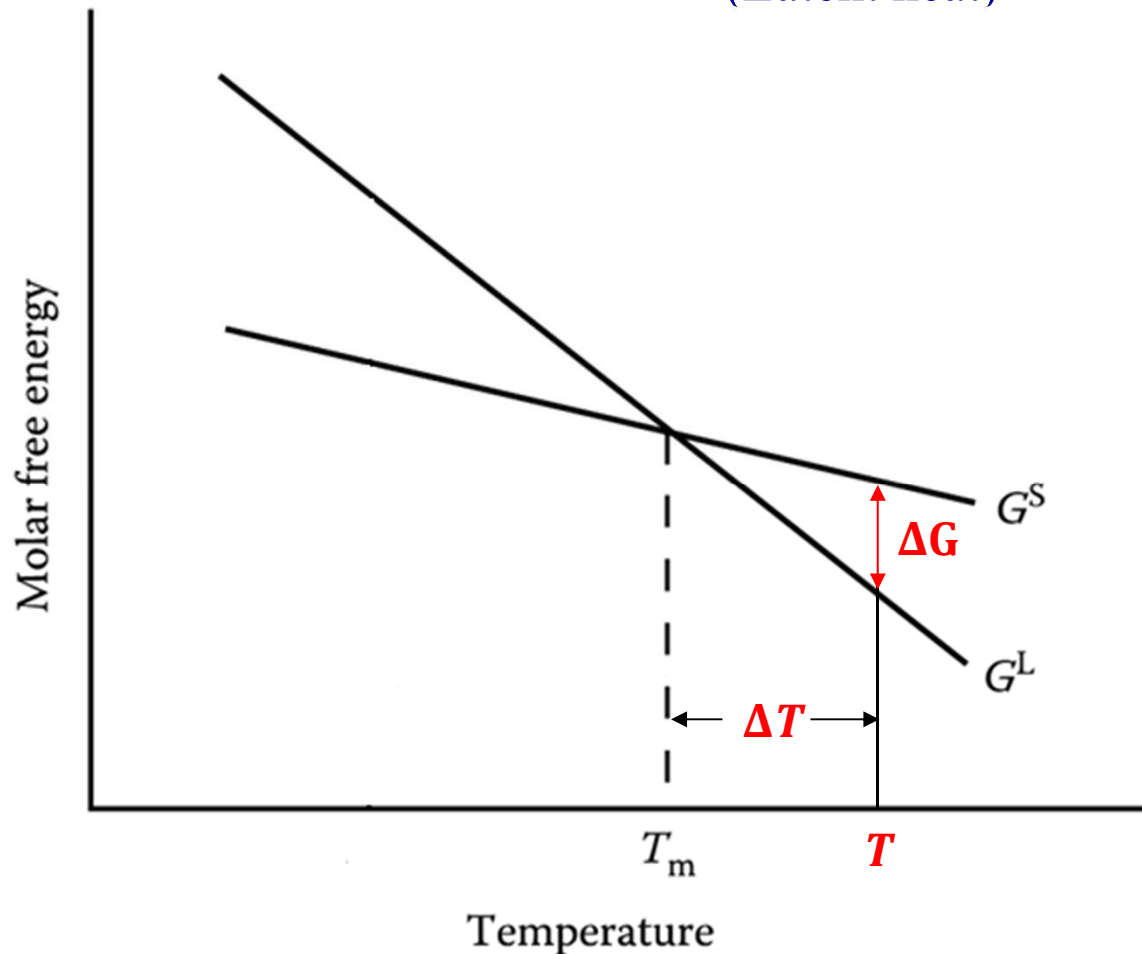
(Latent heat)

$$\Delta G = 0 = \Delta H - T_m \Delta S$$

$$\Delta S = \Delta H / T_m = L / T_m$$

$$\Delta G = L - T(L/T_m) \approx (L\Delta T) / T_m$$

(eq. 1.17)



$$\Delta G = \frac{L\Delta T}{T_m}$$

Nucleation of melting

Although nucleation during solidification usually requires some undercooling, melting invariably occurs at the equilibrium melting temperature even at relatively high rates of heating.

Why?

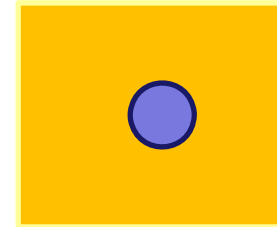
$$\gamma_{SL} + \gamma_{LV} < \gamma_{SV} \quad \text{(commonly)}$$



In general, wetting angle = 0 \rightarrow No superheating required!

Melting and Crystallization are Thermodynamic Transitions

Solidification: Liquid \rightarrow Solid



<Thermodynamic>

• Interfacial energy $\Rightarrow \Delta T_N$

Liquid

T_m

Undercooled Liquid

Solid

No superheating required!

• Interfacial energy \Rightarrow No ΔT_N

$$\gamma_{SL} + \gamma_{LV} < \gamma_{SV}$$

vapor



Melting: Liquid \leftarrow Solid

Contents for today's class

Chapter 1

Thermodynamics and Phase Diagrams

- **Equilibrium** $dG = 0$ **Phase Transformation**
 $\Delta G = G_2 - G_1 < 0$
Lowest possible value of G
No desire to change ad infinitum

- **Single component system**

Gibbs Free Energy as a Function of **Temp.** and **Pressure**

$$\left(\frac{\partial G}{\partial T}\right)_P = -S, \quad \left(\frac{\partial G}{\partial P}\right)_T = V$$
$$\left(\frac{dP}{dT}\right)_{eq} = \frac{\Delta H}{T_{eq} \Delta V} \quad \text{Clausius-Clapeyron Relation}$$

- **Classification of phase transition**

First order transition: **CDD**/Second order transition: **CCD**

- **Driving force for solidification** $\Delta G = \frac{L\Delta T}{T_m}$