Stoichiometry of Biochemical Reactions I

Today's lecture

- Biochemical reaction stoichiometry
- Cell yield
- Half reactions

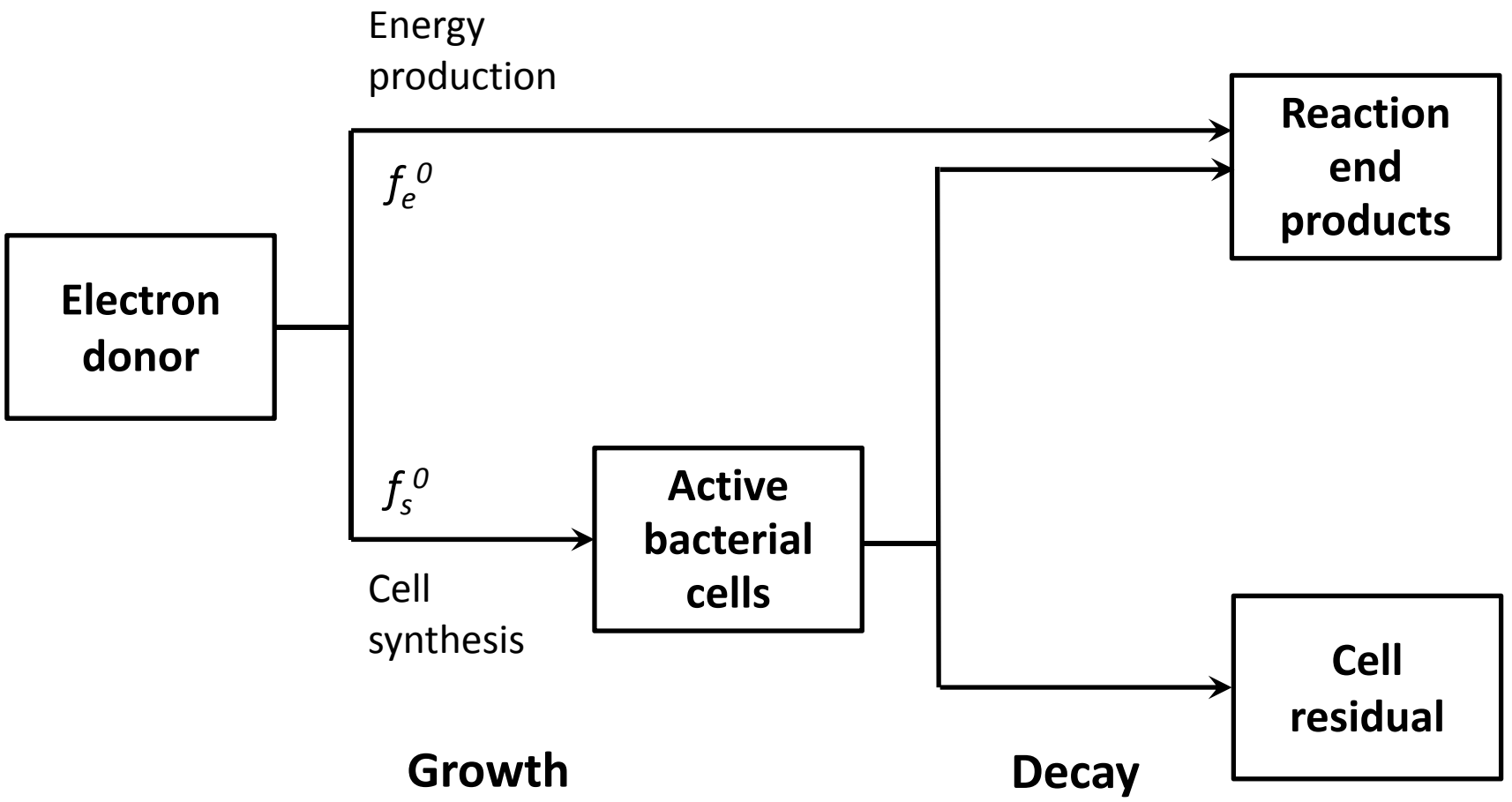
Stoichiometry

- “An aspect of chemistry concerned with mole relationships among reactants and products”
- Based on mass conservation
- Balancing chemical reactions

Cell formula

- Most common: $C_5H_7O_2N$
- COD for a bacterial cell?

Substrate partitioning



Textbook Fig. 2.1

Cell yield

- True yield, *Y*

  Y = (g cells produced) / (g substrate utilized)

- Conversion of f_s^0 to *Y*:

$$Y = f_s{}^0 \frac{(M_c \ \ g\ cells/mole\ cells)}{(n_e \ \ e^-\ eq/mole\ cells)(8\ g\ COD/e^-\ eq\ donor)}$$

Microbial growth rate

$$\frac{dX_a}{dt} = Y\left(\frac{-dS}{dt}\right) - bX_a$$

growth decay

X_a = active biomass concentration [M/L^3]

S = substrate concentration [M/L^3]

Y = true yield [M/M]

b = decay rate [1/T]

Net yield

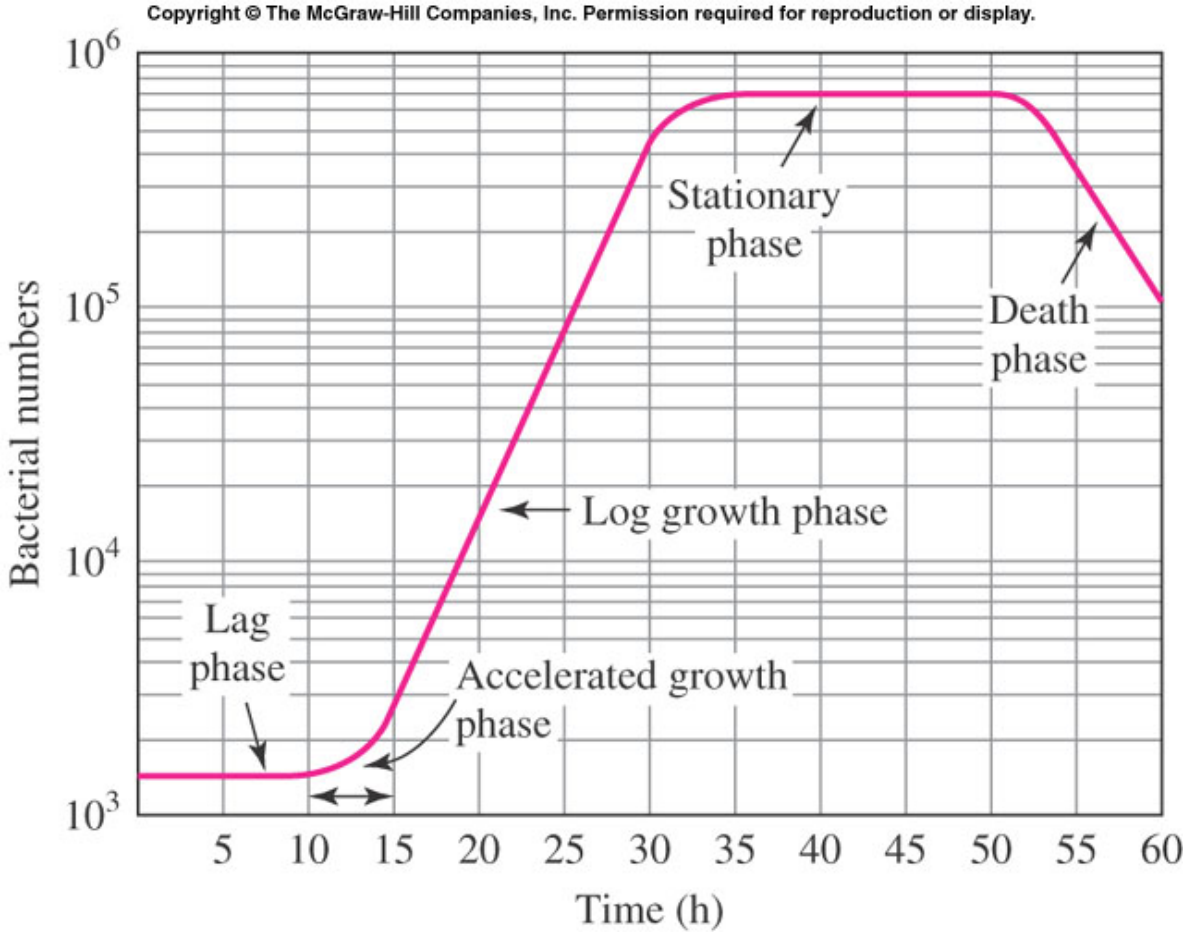
- Net yield, Y_n

  Y_n = (g net cell growth) / (g substrate utilized)

$$= \frac{dX_a / dt}{-dS / dt}$$

$$= Y - b\frac{X_a}{-dS / dt}$$

Net yield



Bacterial growth curve for pure culture

Log (exponential) growth:

Stationary phase:

Death phase:

Net yield

- Electron partitioning considering net yield, Y_n:

  $f_s^0 \rightarrow f_s \qquad (f_s < f_s^0)$

  $f_e^0 \rightarrow f_e \qquad (f_e > f_e^0)$

  still, $f_s + f_e = 1$

Energy reactions

Aerobic oxidation:

$C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$, $\Delta G = -2880$ kJ/mole glucose

vs.

Sulfate reduction:

$2C_6H_{12}O_6 + 6SO_4^{2-} \rightarrow 12CO_2 + 12H_2O + 3H_2S + 3\ HS^-$, $\Delta G = -492$ kJ/mole glucose

Half reactions

- For complex biochemical redox reactions, it is easier to use half reaction approach
- The oxidation reaction for an electron donor and the reduction reaction for an electron acceptor can be splitted
- Usually written as a reduction reaction

  (see [Table 2.2] & [Table 2.3])

Half reactions

Step 1 Write oxidized form on the left and reduced form on the right

Step 2 Add other species involved in the reaction

Step 3 Balance the reaction for all elements except for oxygen and hydrogen

Step 4 Balance oxygen using water

Step 5 Balance hydrogen using H^+

Step 6 Balance charge using e^-

Step 7 Convert the equation to the e^--equivalent form