

Physical Chemistry

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matter

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logic

origin

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change

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study

Equilibrium + Structure + Change

<http://bp.snu.ac.kr>

Introduction

- Mass of proton = $2000 \times$ mass of electron
- Electrons contribute very little to the total mass of the atom, but occupy an appreciable volume, and are responsible for the atomic bulk.

The Quantum Rules:

“The energy of an object cannot be changed by an arbitrary amount.”

Consider a pendulum.

$0, h\nu, 2h\nu, 3h\nu, \dots$

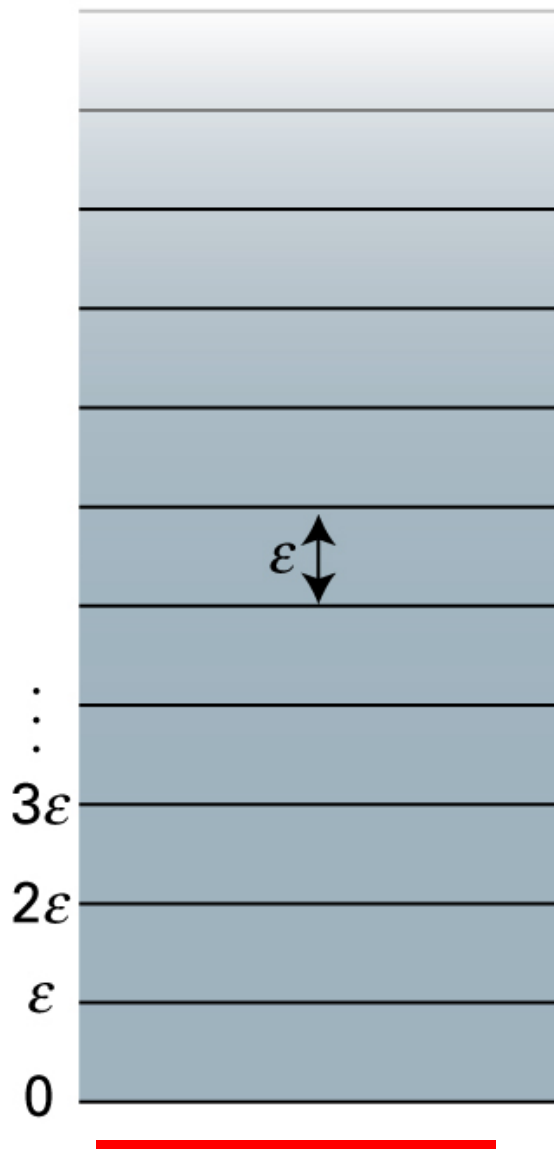
h : Planck's constant

ν : natural frequency

pendulum or spring

$$\nu = \frac{1}{2\pi} \sqrt{\frac{g}{l}} = \frac{1}{2\pi} \sqrt{\frac{k}{m}}$$

Reference Point



Reference Point ?

Figs. 8B.1 and 8B.2

Fig. 15B.11 Harmonic Oscillator

The array of molecular energy levels used to calculate the molecular partition function.

Chap. 15

Phonon = Lattice Vibration

Pendulums of different natural frequencies accept and give up quanta of different sizes, and the higher the natural frequency the larger the quantum energy.

$$h = 4.1357 \times 10^{-15} \text{ eV}\cdot\text{s}$$

if $\nu = 1 \text{ Hz}$, $h\nu = 4.1357 \times 10^{-15} \text{ eV} \longrightarrow$ too small to be detected.

Only when very high-frequency motions are involved, the quantum energy $h\nu$ becomes large enough to give rise to significant effects.

A mass on a spring oscillating with simple harmonic motion: **Phonon**

$$\nu = \frac{1}{2\pi} \sqrt{\frac{k}{m}}$$

$$\nu \approx 10^{13} \text{ Hz}$$

\therefore When the mass is that of an atom, the frequency is extremely high, and so we must be prepared for the significant quantum effects in vibrating molecules.

Photon

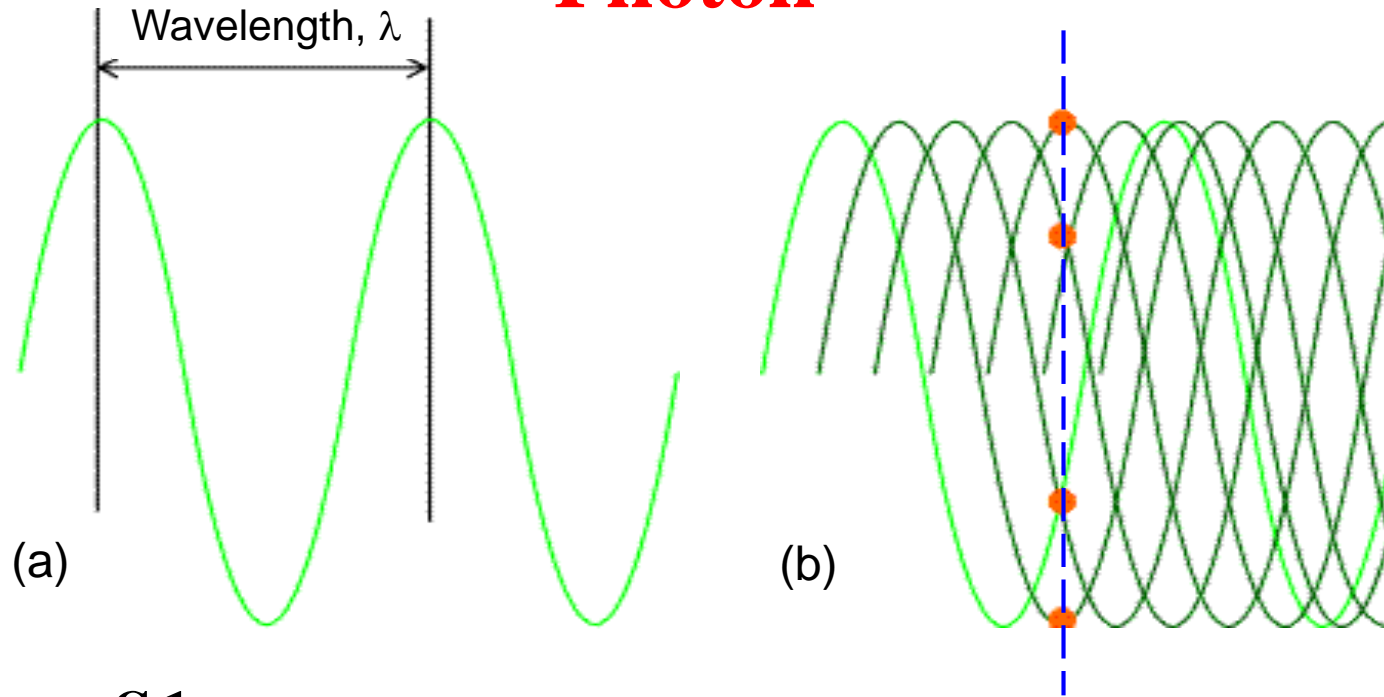


Figure C.1

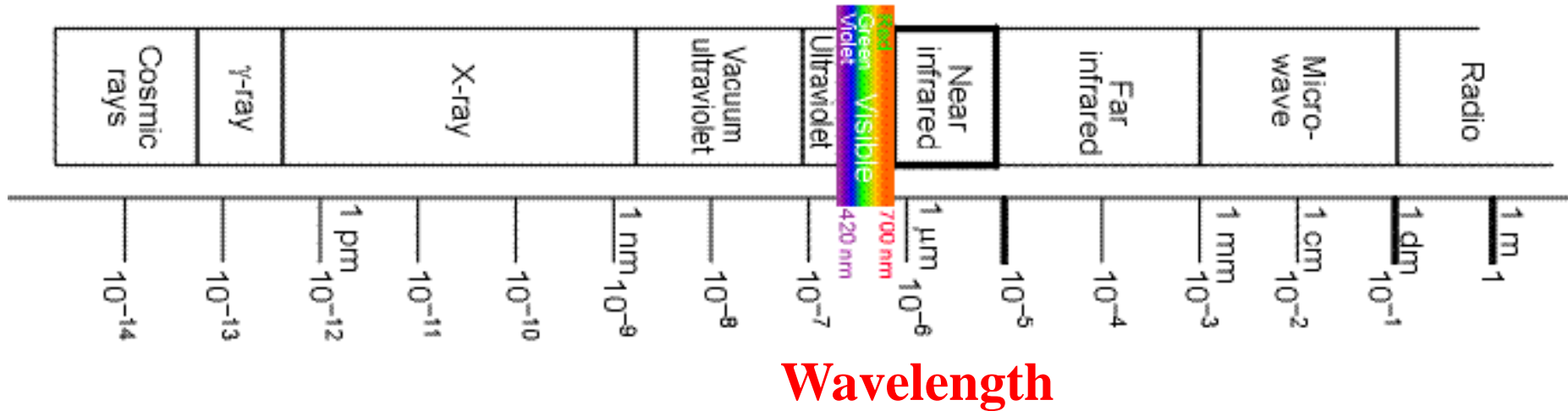
The wave is shown travelling to the right at a speed c .

Wavelength and frequency are related by $\lambda \nu = c$.

Figures C.5 and C.6

Plane-polarized and circularly-polarized waves.

Nuclear excitation Core-electron excitation Electronic excitation Molecular vibration Molecular rotation



Wavelength

Figure C.4

The regions of the electromagnetic spectrum and the types of excitation that give rise to each region.

Visible Light: $\sim 400 - 700$ nm

$\sim 2 - 3$ eV

$\sim 10^{15}$ Hz

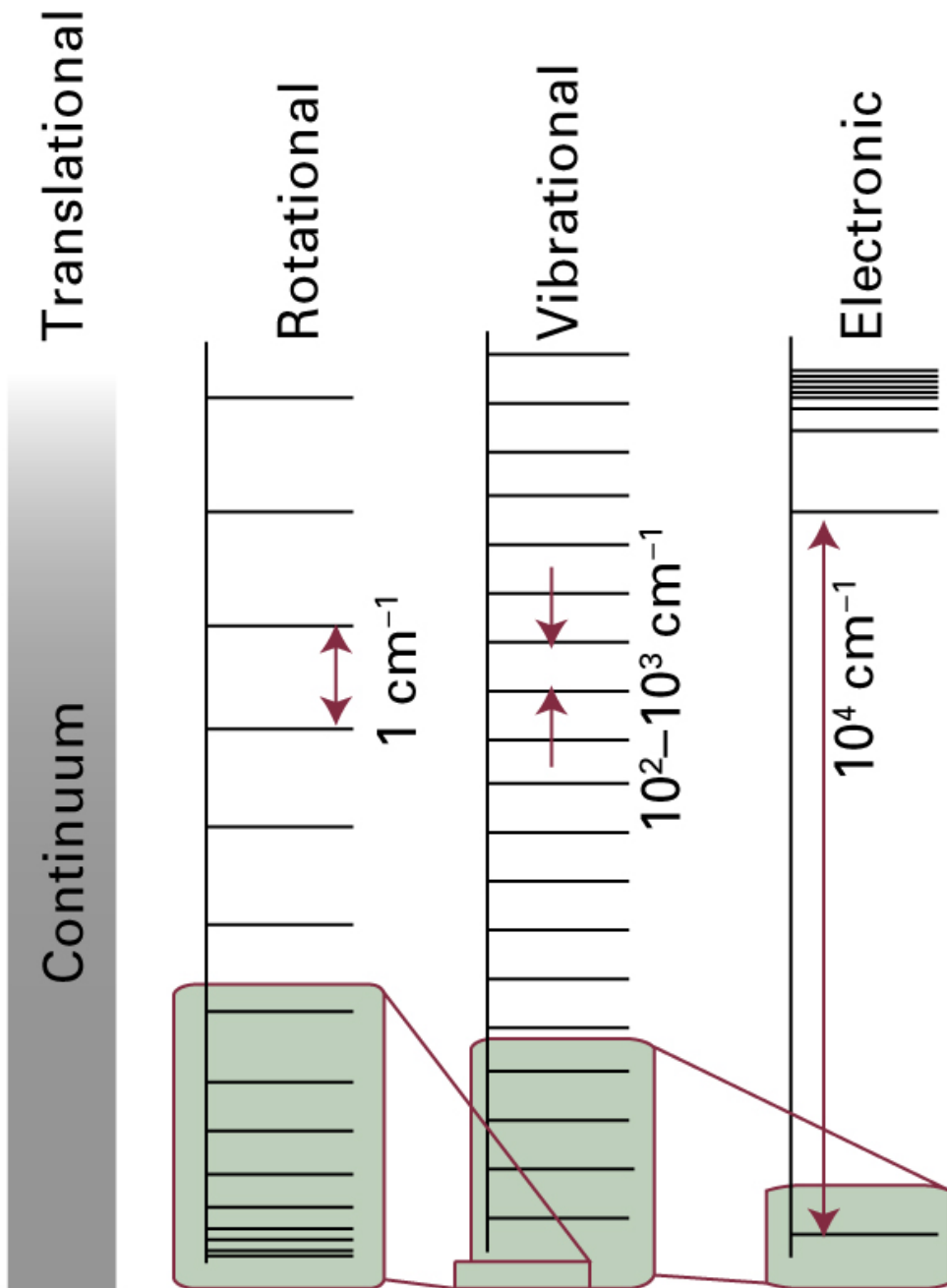


Figure B.3

Quantization of the energy of different types of motion.

Free **translational** motion in an infinite region looks not quantized, and the permitted energy levels form a continuum.

Rotation is quantized, and the separation increases as the state of excitation increases (moment of inertia of the molecule).

Vibrational motion is quantized, but note the change in scale between the ladders.

Electronic energy levels are quantized, and the separations are typically very large (of the order of 1 eV).

B.3(a) Boltzmann Distribution

*Derivation at
Chapter 15A*

Consider **Thermal Agitation**.

- Some molecules occupy the lowest energy levels, others will be in the excited levels.
- Not all reach the higher levels: there is a predominance of molecules in the lower, less energetic states, but the tail of the distribution gets larger and penetrates further into the high energy region as the temperature is raised (Fig. F.5).

$$\frac{N_i}{N_j} = \exp[-(E_i - E_j)/RT]$$

(B.25a)

(B.25b)

E_i, E_j : **molar** energies at states i and j

R : gas constant, 8.314 JK^{-1}

T : absolute temperature **Boltzmann's Constant:**

$$k_B = 8.617 \times 10^{-5} \text{ eV/K}$$

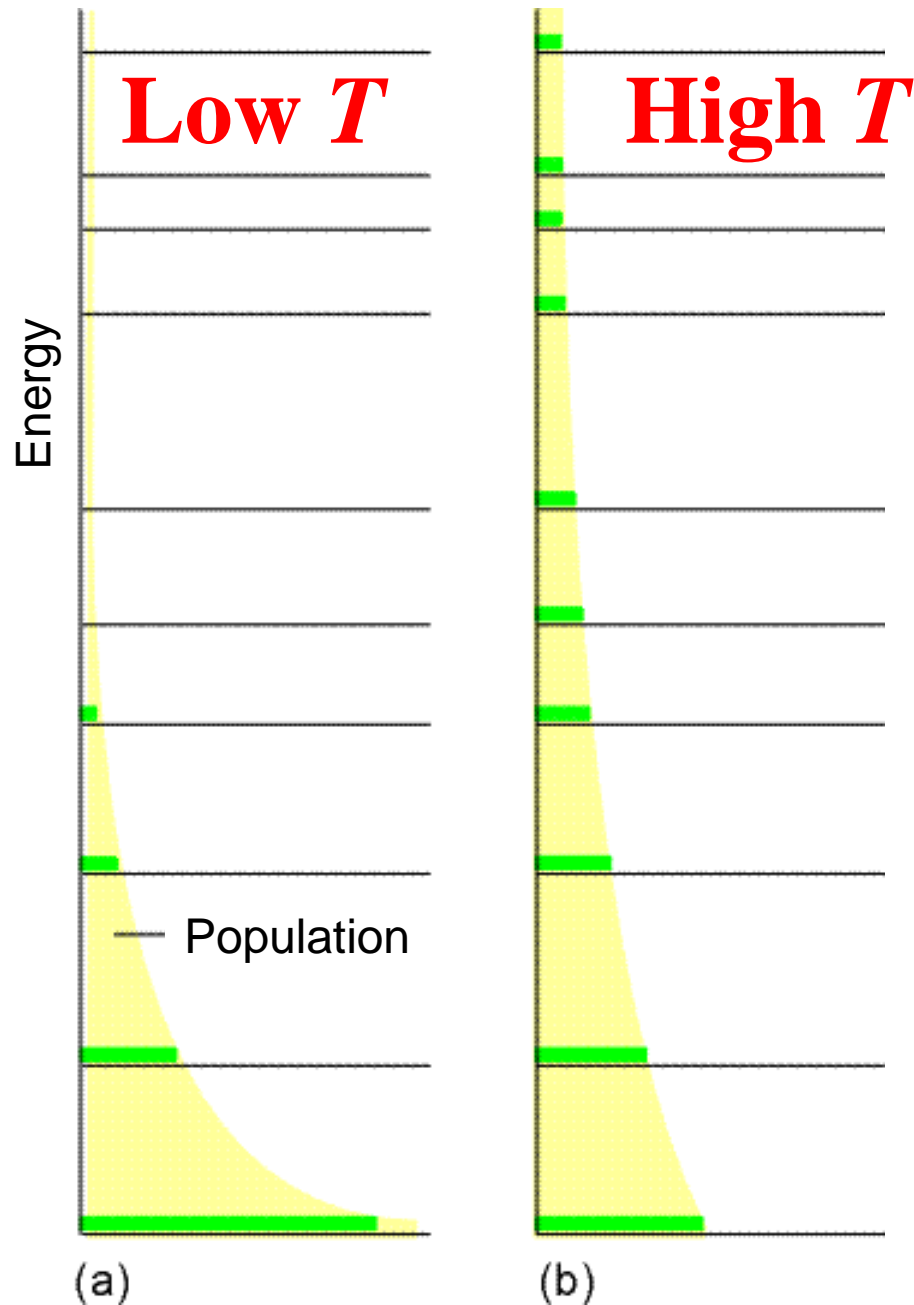


Figure B.4

The Boltzmann distribution predicts that the population of a state decreases exponentially with the energy of the state.

- (a) At low temperatures, only the lowest states are significantly populated.
- (b) At high temperatures, there is significant population in high-energy states as well as in low-energy states. At infinite temperature (not shown), all states are equally populated.

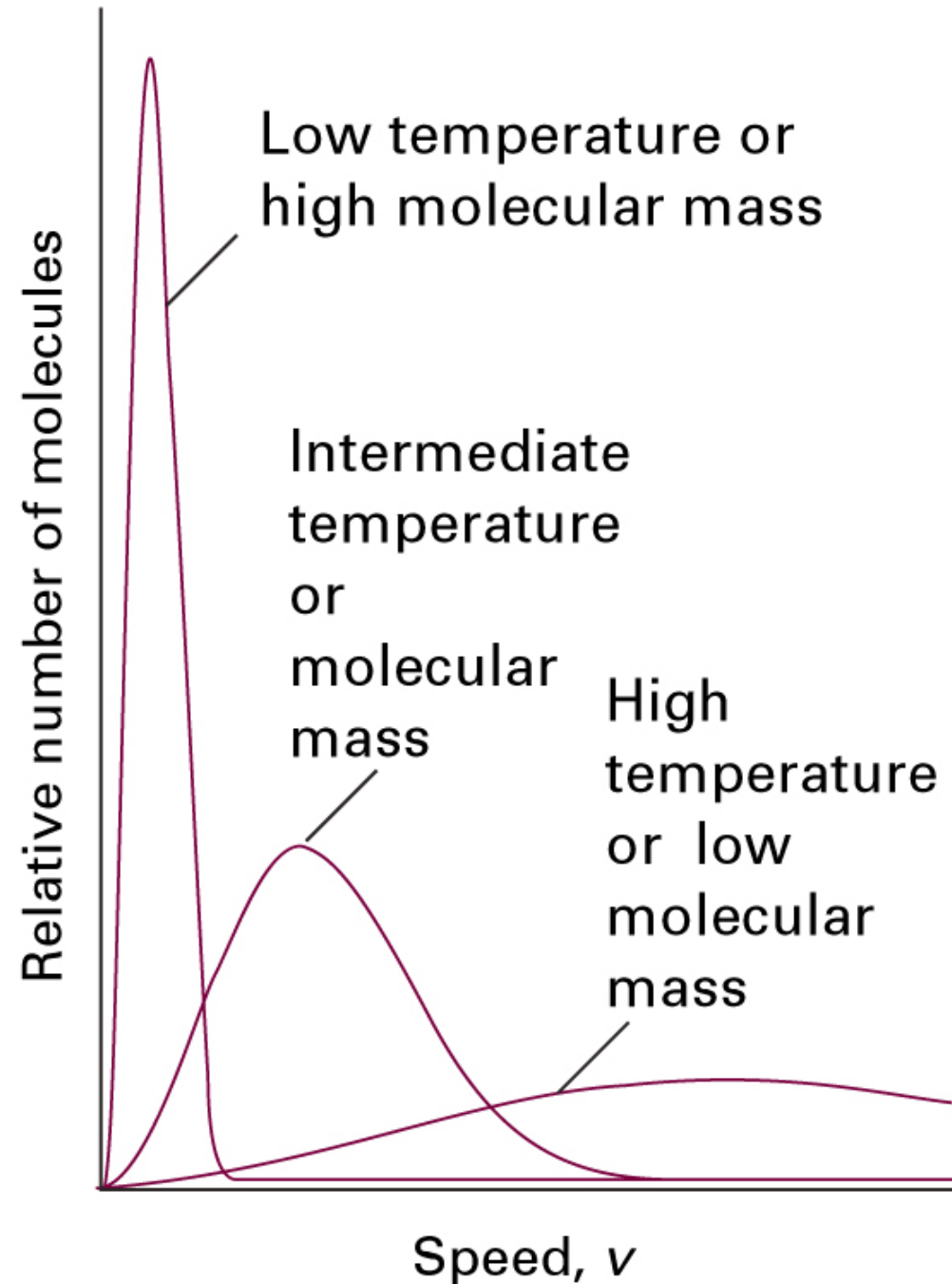


Figure B.6

The essential content of the

Maxwell-Boltzmann
distribution of
molecular speeds.

Note how the maximum in the distribution moves to higher speeds as the temperature is increased or, at constant temperature, we consider species of decreasing mass.

The distribution also becomes wider as its peak moves to higher speeds.

The average kinetic energy of 1 mol of particles from the Maxwell-Boltzmann distribution = $\frac{3}{2}RT$.

B.3(b)
Chapter 15C

Equipartition theorem of classical physics:

average energy of different modes of motion = $\frac{1}{2}RT$.

\therefore three dimensional translational modes = $3 \times \frac{1}{2}RT$

Association of Atoms and Molecules

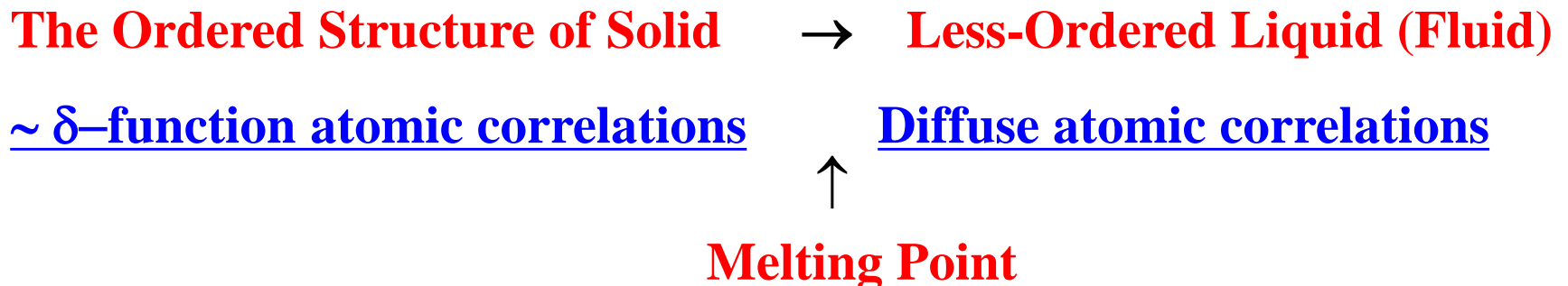
1. Ionic bond
2. Covalent bond
3. Metallic bond: an assembly of cations (Cu^{2+}) immersed in a sea of electrons
4. van der Waals force:
 - (a) due to polar attraction
 - (b) due to instantaneous electric dipoles
 - (c) due to **hydrogen bonding**

A.3 The State of Matter

Solid State: **Crystal** or **Amorphous (Glass)**

1. Morphology: crystal's external appearance
2. Crystal defects: the breaks in the uniformity of array
3. Liquid crystals: van der Waals attraction is so weak that the crystals possess virtually no rigidity.

Solid State vs. Liquid State



The Ability to Flow

The number of molecules with enough energy to move follows a **Boltzmann distribution**.

$$\text{Fluidity} \propto \exp(-\Delta E_a / RT)$$

where ΔE is the energy to be overcome.

$$\text{Viscosity} \propto \exp(+\Delta E / RT)$$

where ΔE is of the same order as the intermolecular binding energy.

Activation Energy

Activation Enthalpy

Gas Phase

The average distance between molecules is normally large relative to their diameter. At low pressure, the intermolecular interactions are likely to be negligible, and the gas satisfies **the ideal gas (perfect gas) law**:

$$pV = nRT \quad p: \text{pressure} \quad (\text{A.5})$$

n : amount of gas

T : absolute temperature

V : volume

The mean kinetic energy of 1 mol of gas $= \frac{3}{2} RT$

The kinetic energy of 1 mol of gas $= \frac{1}{2} m \langle v^2 \rangle N_A$

N_A : Avogadro's number

$\langle v^2 \rangle$: average value of v^2

$$\frac{3}{2} RT = \frac{1}{2} m \langle v^2 \rangle N_A$$

$$\underline{\frac{3}{2} k_B T = \frac{1}{2} m \langle v^2 \rangle}$$

$$\langle v^2 \rangle = \frac{3RT}{mN_A}$$

$$\sqrt{\langle v^2 \rangle} = \sqrt{\frac{3RT}{mN_A}} = \sqrt{\frac{3k_B T}{m}} \quad : \text{Root mean square speed}$$

$$k_B = \frac{R}{N_A}$$

Transformations of Matter

- Can it occur? **Thermodynamics**
- How fast does it occur? **Kinetics**

The probability that, when the collision occurs, it does so with sufficient energy to react will be determined by **the Boltzmann distribution**.

E_a : necessary energy (**activation energy** of the reaction)

Rate of reaction \propto (rate of collision) \times (probability that
collision carries
enough energy)

$$\propto Z(T) \exp\left(-\frac{E_a}{RT}\right) \quad \textbf{Arrhenius equation}$$

Chapter 20D

Table F.1 The SI base units

Physical quantity	Symbol for quantity	Base unit
Length	l	metre, m
Mass	m	kilogram, kg
Time	t	second, s
Electric current	I	ampere, A
Thermodynamic temperature	T	kelvin, K
Amount of substance	n	mole, mol
Luminous intensity	I_v	candela, cd

cm g s A K mol cd

Table F.2 A selection of derived units

Physical quantity	Derived unit*	Name of derived unit
Force	1 kg m s^{-2}	newton, N
Pressure	$1 \text{ kg m}^{-1} \text{ s}^{-2}$ 1 N m^{-2}	pascal, Pa
Energy	$1 \text{ kg m}^2 \text{ s}^{-2}$ 1 N m 1 Pa m^3	joule, J
Power	$1 \text{ kg m}^2 \text{ s}^{-3}$ 1 J s^{-1}	watt, W

* Equivalent definitions in terms of derived units are given following the definition in terms of base units.

$$1 \text{ eV} = 1.602 \times 10^{-19} \text{ J} = 1.602 \times 10^{-12} \text{ erg}$$

Table F.3 Common SI prefixes

Prefix	y	z	a	<u>f</u>	p	n	μ	m	c	d
Name	yocto	zepto	atto	femto	pico	nano	micro	milli	centi	deci
Factor	10 ⁻²⁴	10 ⁻²¹	10 ⁻¹⁸	10 ⁻¹⁵	10 ⁻¹²	10 ⁻⁹	10 ⁻⁶	10 ⁻³	10 ⁻²	10 ⁻¹
Prefix	da	h	<u>k</u>	M	G	<u>T</u>	P	E	Z	Y
Name	deca	hecto	kilo	mega	giga	tera	peta	exa	zeta	yotta
Factor	10	10 ²	10 ³	10 ⁶	10 ⁹	10 ¹²	10 ¹⁵	10 ¹⁸	10 ²¹	10 ²⁴

Table F.4 Some common units

Physical quantity	Name of unit	Symbol for unit	Value*
Time	minute	min	60 s
	hour	h	3600 s
	day	d	86 400 s
	year	a	31 556 952 s
Length	ångström	Å	10^{-10} m
Volume	litre	L, l	1 dm^3
Mass	tonne	t	10^3 kg
Pressure	bar	bar	10^5 Pa
	atmosphere	atm	101.325 kPa
Energy	electronvolt	<u>eV / atom</u>	$1.602\,176\,53 \times 10^{-19}$ J
		$\approx 10^5$ J/mol	<u>96.485 31 kJ mol⁻¹</u>

* All values in the final column are exact, except for the definition of 1 eV, which depends on the measured value of *e*, and the year, which is not a constant and depends on a variety of astronomical assumptions.

eV / molecule

Problems from Chap. Foundations

A.1(b) A.2(b) A.3(b) A.10(b) A.19(b)

B.2 B.4 B.6

B.4(b) B.8(b) B.9(b) B.14(b) B.15(b)

B.17(b) B.19(b) B.20(b) B.23(b) B.24(b)

C.1(b) C.2(b)